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4624Tandem templating strategies facilitate the  
assembly of calix[8]arene-supported Ln<sub>18</sub> clusters†Yushu Jiao,<sup>a</sup> Sergio Sanz,<sup>b</sup> Jan van Leusen,<sup>a</sup> David Gracia,<sup>c</sup>  
Angelos B. Canaj,<sup>d</sup> Marco Evangelisti,<sup>e</sup> Euan K. Brechin,<sup>d</sup>  
Scott J. Dalgarno<sup>e</sup> and Paul Kögerler<sup>a,b</sup>

Calix[*n*]arenes offer ideal chemical functionality through the polyphenolic lower rim to construct nano-sized coordination clusters with lanthanide (Ln) metal ions (e.g., Nd<sup>III</sup>, Gd<sup>III</sup>). However, the number of metal centers they can accommodate is still limited compared to that achievable with smaller ligands (e.g., Gd<sup>III</sup><sub>140</sub>, Gd<sup>III</sup><sub>104</sub>). Here, we exploit a combination of the “anion template strategy” and “templating ligands” to synthesise three highly symmetric (*D*<sub>3h</sub>, trigonal planar) Ln<sup>III</sup><sub>18</sub> (Ln = La, Nd, and Gd) systems, representing the largest calix[*n*]arene-based coordination clusters yet. The Ln<sup>III</sup><sub>18</sub> fragment is templated by a chloride anion located at the center of the cluster, wherefrom twelve μ<sub>3</sub>-OH<sup>-</sup> ligands bind ‘internally’ to the eighteen Ln<sup>III</sup> ions. ‘Externally’ the metallic skeleton is connected by *p*-*tert*-butylcalix[8]arene, oxo, chloro and carbonate ligands. The crystal packing in the lattice reveals large cylindrical channels of ~26 Å in diameter, whose pore volume corresponds to ~50% of the unit cell volume (using a 1.2 Å spherical probe radius). Magnetic measurements reveal the predominance of weak antiferromagnetic exchange in the Gd analog. Heat capacity data of Gd<sup>III</sup><sub>18</sub> reveal a high magnetic entropy with  $-\Delta S_m = 23.7 \text{ J K}^{-1} \text{ kg}^{-1}$ , indicating potential for engineering magnetic refrigerant materials with calix[8]arenes.

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## Introduction

The synthesis of high nuclearity, homometallic lanthanide (Ln) coordination clusters has been developed in recent decades owing to their fascinating topological configurations and unique chemical and physical properties, with potential applications in fields as diverse as catalysis, optics, and magnetism.<sup>1</sup> The synthesis of such clusters entails ligands that are able to define the outer sphere of the molecule by limiting the uncontrollable hydrolysis of the metal ions. The latter is characterized by the production of intractable precipitates of

metal hydroxides and/or oxohydroxides. Several classes of ligands have been successfully tested for this purpose, producing giant homometallic Ln clusters including Gd<sup>III</sup><sub>140</sub>,<sup>1a</sup> Ln<sup>III</sup><sub>104</sub>,<sup>1b</sup> Dy<sup>III</sup><sub>76</sub>,<sup>1c</sup> Dy<sup>III</sup><sub>72</sub>,<sup>1d</sup> Ln<sup>III</sup><sub>60</sub>,<sup>1e,f</sup> and Ln<sup>III</sup><sub>48</sub>,<sup>2</sup> whose metallic cores are commonly constructed by the repetition of a small number of well-defined building blocks constructed from metal-hydroxo motifs. These include triangular {Ln<sub>3</sub>(μ<sub>3</sub>-OH)}, tetrahedral {Ln<sub>4</sub>(μ<sub>3</sub>-OH)<sub>4</sub>} and trigonal bipyramidal {Ln<sub>5</sub>(μ<sub>3</sub>-OH)<sub>6</sub>}. The metallic skeletons of these clusters therefore tend to conform to specific polyhedra possessing the same characteristic symmetries as their building blocks (e.g., tetrahedral *T*<sub>d</sub>, cubic *O*<sub>h</sub>, dodecahedron *I*<sub>h</sub>, etc.).<sup>3</sup>

Perhaps the most successful calix[*n*]arene ligands in the design and construction of high-nuclearity metal coordination clusters are the *p*-*tert*-butylcalix[*n*]arenes (H<sub>*n*</sub>TBC[*n*]) comprising either 4 or 8 phenolic units (H<sub>4</sub>TBC[4] and H<sub>8</sub>TBC[8]). Although both have been commonly employed in supramolecular chemistry (3331 and 161 hits in the Cambridge Structural Database, CSD, respectively), their use as coordination ligands to build homometallic Ln clusters remains relatively underexplored (16 and 15 hits in the CSD for H<sub>4</sub>TBC[4] and H<sub>8</sub>TBC[8], respectively). The vast majority are dimeric Ln<sub>2</sub> species, particularly for *p*-*tert*-butylcalix[8]arene. Examples of larger nuclearity cages include [Nd<sup>III</sup><sub>10</sub>(TBC[8])<sub>2</sub>], [Gd<sup>III</sup><sub>8</sub>(TBC[8])<sub>2</sub>], [Dy<sup>III</sup><sub>7</sub>(TBC[8])<sub>2</sub>], and [Tb<sup>III</sup><sub>6</sub>(TBC[4])<sub>2</sub>].<sup>4</sup>

<sup>a</sup>Institute of Inorganic Chemistry, RWTH Aachen University, 52056 Aachen, Germany. E-mail: paul.koegerler@ac.rwth-aachen.de

<sup>b</sup>Peter Grünberg Institute, Electronic Properties (PGI-6), Forschungszentrum Jülich, 52425 Jülich, Germany. E-mail: s.calvo@fz-juelich.de

<sup>c</sup>Instituto de Nanociencia y Materiales de Aragón (INMA), CSIC – Universidad de Zaragoza, Departamento de Física de la Materia Condensada, 50009 Zaragoza, Spain. E-mail: evange@unizar.es

<sup>d</sup>EaStCHEM School of Chemistry, The University of Edinburgh, David Brewster Road, Edinburgh, EH9 3FJ, UK. E-mail: ebrechin@ed.ac.uk

<sup>e</sup>Institute of Chemical Sciences, Heriot-Watt University, Riccarton, Edinburgh, EH14 4AS, UK. E-mail: S.J.Dalgarno@hw.ac.uk

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The synthesis of polynuclear coordination clusters with a particular blend of physical properties has resulted in magnetic molecules that can potentially be employed as efficient cryogenic magnetic refrigerants *via* the magnetocaloric effect (MCE).<sup>5</sup> Magnetic entropy ( $\Delta S_m$ ) and adiabatic temperature ( $\Delta T_{ad}$ ) changes take place in a magnetocaloric material following the change of an applied magnetic field ( $\Delta B$ ). The ideal metal ion to exploit is Gd<sup>III</sup> due to its large isotropic spin  $s = 7/2$ , leading to several high-nuclearity coordination clusters that possess large molecular magnetic entropies at low temperatures.<sup>6</sup> Such species are energy-efficient alternatives to traditional cooling materials such as <sup>3</sup>He or <sup>4</sup>He.<sup>7</sup>

We have recently focused at a breadth of different synthetic conditions in calix[*n*]arene coordination chemistry, including the implementation of the “anion template strategy” and the utilization of “templating ligands”. The former has long proven adept in constructing high-nuclearity clusters with a range of organic ligands, but strangely has been almost neglected in calix[*n*]arene coordination chemistry. The introduction of “templating ligands” – critical for the molecule’s formation but not integrated into its structure – is believed to play a crucial role in directing the self-assembly of large nuclearity species by isolating the smaller building blocks, preventing rapid aggregation into smaller complexes with more stable topologies. The recent successful synthesis of the molecular iron oxides Fe<sup>III</sup><sub>34</sub> and Fe<sup>III</sup><sub>30</sub>,<sup>8</sup> achieved using this methodology, provides an excellent demonstration of its potential and relevance to be introduced in the field of calixarene chemistry.

Here we adopt these strategies in the synthesis of three highly symmetric (*D*<sub>3h</sub>, trigonal planar) Ln<sup>III</sup><sub>18</sub> clusters of formula [Ln<sup>III</sup><sub>18</sub>(TBC[8])<sub>3</sub>(μ<sub>4</sub>-O)<sub>3</sub>(μ<sub>3</sub>-Cl)<sub>6</sub>(μ<sub>3</sub>-OH)<sub>12</sub>(μ<sub>3</sub>-CO<sub>3</sub>)<sub>2</sub>(H<sub>2</sub>O)<sub>6</sub>(DMF)<sub>18</sub>Cl]OH·H<sub>2</sub>O·12DMF, representing the largest known calix[*n*]arene-based coordination clusters.

## Experimental section

### Synthetic procedure

All solvents and reagents (99.9%) used in this work were obtained from commercial sources. [Ln<sup>III</sup><sub>2</sub>(H<sub>2</sub>TBC[8]) (DMSO)<sub>3</sub>(H<sub>2</sub>O)]·2H<sub>2</sub>O·3DMSO was synthesized according to a modified literature<sup>9</sup> method as described below.

### Preparation of [Ln<sup>III</sup><sub>2</sub>(H<sub>2</sub>TBC[8]) (DMSO)<sub>3</sub>(H<sub>2</sub>O)]·2H<sub>2</sub>O·3DMSO (Ln<sub>2</sub>)

A mixture of H<sub>8</sub>TBC[8] (150 mg, 0.115 mmol), Ln<sup>III</sup>Cl<sub>3</sub>·6H<sub>2</sub>O (132.51 mg for La<sup>III</sup>Cl<sub>3</sub>·6H<sub>2</sub>O, 134.50 mg for Nd<sup>III</sup>Cl<sub>3</sub>·6H<sub>2</sub>O, and 139.38 mg for Gd<sup>III</sup>Cl<sub>3</sub>·6H<sub>2</sub>O; 0.375 mmol) in 14 mL DMSO/MeCN (*v*:*v* = 1:1) was stirred for 5 min and then triethylamine (0.18 mL, 1.29 mmol) was added. The solution was further stirred for another 2 h. The resulting solution was left to stand at room temperature overnight to allow the formation of a white precipitate. The crude product was then isolated by filtration and purified by slow diffusion of *tert*-butyl methyl ether into a CHCl<sub>3</sub> solution of the product affording colorless needle-shape crystals of [Ln<sup>III</sup><sub>2</sub>(H<sub>2</sub>TBC[8]) (DMSO)<sub>3</sub>(H<sub>2</sub>O)]·2H<sub>2</sub>O·3DMSO.

### Synthesis of [La<sup>III</sup><sub>18</sub>(TBC[8])<sub>3</sub>(μ<sub>4</sub>-O)<sub>3</sub>(μ<sub>3</sub>-Cl)<sub>6</sub>(μ<sub>3</sub>-OH)<sub>12</sub>(μ<sub>3</sub>-CO<sub>3</sub>)<sub>2</sub>(H<sub>2</sub>O)<sub>6</sub> (DMF)<sub>18</sub>Cl]OH·H<sub>2</sub>O·12DMF (1)

A solution of La<sub>2</sub> (314.45 mg, 0.15 mmol) and 1,3,5-tricarboxylic acid (15.75 mg, 0.075 mmol) in 8 mL CHCl<sub>3</sub>/DMF/MeOH (*v*:*v*:*v* = 3:3:2) was stirred for 5 min. Afterwards, NBu<sub>4</sub>OH (0.1 mL, 40% in water) was slowly added to the solution under stirring and left to stir for 1 hour. The resulting mixture was then sealed in a PTFE-lined bomb and heated to 120 °C for 48 h. After cooling slowly to room temperature, the cloudy solution was filtered to yield a clear yellow solution. Brown, block-shaped crystals were obtained by slow diffusion of hexane into the mother liquor over a period of four weeks. Yield (5 mg, 3.6%). IR ( $\tilde{\nu}/\text{cm}^{-1}$ ): 2953 (m), 2903 (w), 2867 (w), 1653 (vs), 1598 (w), 1453 (s), 1390 (m), 1360 (m), 1297 (m), 1267 (m), 1207 (s), 1108 (w), 1023 (m), 907 (w), 862 (w), 818 (m), 804 (w), 739 (m), 679 (m), 606 (w). **Crystal data** (CCDC 2299004): C<sub>356</sub>H<sub>549</sub>Cl<sub>7</sub>La<sub>18</sub>N<sub>30</sub>O<sub>83</sub>, *M* = 9325.75 g mol<sup>-1</sup>, trigonal, space group *P* $\bar{3}1c$  (no. 163), *a* = 32.5846(4) Å, *c* = 34.4461(6) Å, *V* = 31673.5(10) Å<sup>3</sup>, *Z* = 2, *T* = 100(2) K,  $\mu(\text{CuK}\alpha)$  = 9.799 mm<sup>-1</sup>, *D*<sub>calc</sub> = 0.978 g cm<sup>-3</sup>, 458 160 reflections measured (6.264° ≤ 2θ ≤ 130.688°), 18 091 unique (*R*<sub>int</sub> = 0.1166, *R*<sub>sigma</sub> = 0.0431) which were used in all calculations. The final *R*<sub>1</sub> was 0.0832 (*I* > 2σ(*I*)) and *wR*<sub>2</sub> was 0.3071 (all data).

### Synthesis of [Nd<sup>III</sup><sub>18</sub>(TBC[8])<sub>3</sub>(μ<sub>4</sub>-O)<sub>3</sub>(μ<sub>3</sub>-Cl)<sub>6</sub>(μ<sub>3</sub>-OH)<sub>12</sub>(μ<sub>3</sub>-CO<sub>3</sub>)<sub>2</sub>(H<sub>2</sub>O)<sub>6</sub> (DMF)<sub>18</sub>Cl]OH·H<sub>2</sub>O·12DMF (2)

Compound 2 was prepared using the same synthetic procedure as compound 1, with the substitution of La<sub>2</sub> for Nd<sub>2</sub> (134.50 mg, 0.375 mmol). The reaction yielded 6 mg (4.2%) of dark-green block-like crystals. IR ( $\tilde{\nu}/\text{cm}^{-1}$ ): 2951 (m), 2903 (w), 2863 (w), 1653 (vs), 1605 (w), 1460 (s), 1390 (m), 1360 (m), 1296 (m), 1267 (m), 1206 (s), 1107 (w), 1023 (w), 905 (w), 867 (w), 818 (m), 802 (w), 738 (m), 678 (m), 605 (w).

### Synthesis of [Gd<sup>III</sup><sub>18</sub>(TBC[8])<sub>3</sub>(μ<sub>4</sub>-O)<sub>3</sub>(μ<sub>3</sub>-Cl)<sub>6</sub>(μ<sub>3</sub>-OH)<sub>12</sub>(μ<sub>3</sub>-CO<sub>3</sub>)<sub>2</sub>(H<sub>2</sub>O)<sub>6</sub> (DMF)<sub>18</sub>Cl]OH·H<sub>2</sub>O·12DMF (3)

Compound 3 was prepared using the same synthetic procedure as compound 1, replacing La<sub>2</sub> with Gd<sub>2</sub> (139.38 mg, 0.375 mmol) and adjusting the temperature from 120 to 150 °C for 24 h. Brown block-like crystals were obtained by slow evaporation of the mother liquor over four weeks. Yield (4 mg, 2.7%). IR ( $\tilde{\nu}/\text{cm}^{-1}$ ): 2953 (m), 2905 (w), 2865 (w), 1655 (vs), 1611 (w), 1469 (s), 1390 (m), 1360 (m), 1299 (m), 1266 (m), 1207 (s), 1110 (w), 1025 (w), 908 (w), 864 (w), 819 (m), 806 (w), 741 (m), 682 (m), 606 (w).

### Crystal structure information

A single crystal of 1 was selected and measured on a Bruker D8 Venture diffractometer operating with Cu-Kα radiation ( $\lambda = 1.54178$ ) at 100(2) K. Using Olex2,<sup>10</sup> the structure was solved with the SHELXT<sup>12</sup> structure solution program using Intrinsic Phasing and refined with the SHELXL<sup>11</sup> refinement package using Least Squares minimization. There is a considerable amount of disorder present in the structure, which is not



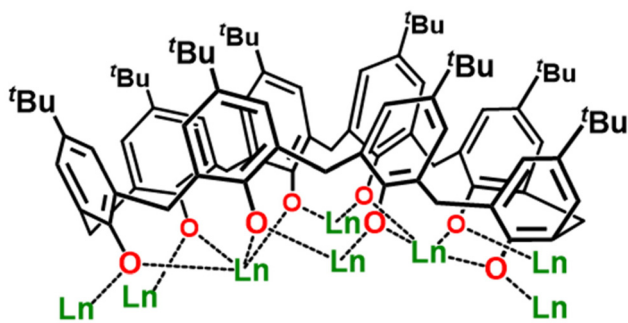


Fig. 1 Representation of *p*-*tert*-butylcalix[8]arene (TBC[8]) and its coordination abilities presented in  $\text{Ln}^{\text{III}}$ . The ligand is in a twofold double-cone conformation, adopting a  $\mu_2$ -O binding mode in each phenoxide unit toward the Ln ions.

uncommon for one of this nature, particularly when channels are formed by pseudo-spherical assemblies that necessarily pack to form solvent-occupied channels or regions in the crystal lattice. A number of the upper-rim *t*Bu groups of the calix[8]arene are disordered and have been modeled at partial occupancies. The hydroxide counterion and waters of crystallization were also handled in this way, whilst co-crystallized DMF was dealt with *via* the use of a solvent mask as it could not be modeled effectively. Given the level of disorder it was necessary to employ numerous restraints, but considering the severity of the problem it was still possible to arrive at a very reasonable agreement index ( $R_1 = 0.0832$ ). Unit cell checks were performed on the  $\text{Gd}^{\text{III}}$  and  $\text{Nd}^{\text{III}}$  analogues (Table S2<sup>†</sup>), confirming these to be isostructural (see ESI<sup>†</sup>). Crystals of both 2 and 3 diffracted poorly relative to those of 1, therefore only this analogue has been reported in full.

### Magnetic measurements

Magnetic measurements were carried out on a Quantum Design MPMS SQUID magnetometer in direct current (dc) mode. Microcrystalline samples of 2 and 3 were separately compacted and immobilized into cylindrical PTFE sample holders. Experimental dc data were recorded at 0.1 T and 1.0 T in the temperature range 2.0–290 K, and isothermal curves up to 5 T in the range of 2–13 K. All data were corrected for the diamagnetic contributions of the sample holders and the complex. AC susceptibility data were also routinely collected at zero static bias field in the 2.0–50 K/3–1000 Hz range using an amplitude of  $B_{\text{ac}} = 3$  G. However, no relevant out-of-phase signals were detected for both compounds. For the evaluation of the magnetocaloric effects of 3, magnetic susceptibility data were collected on polycrystalline powders on a Quantum Design Dynacool PPMS equipped with a 9 T magnet in the temperature range 300–2 K. The samples were added to Quantum Design gelatin capsules with eicosane present and then transferred to a Quantum Design sample straw (AGC2). Diamagnetic corrections from the holders and eicosane were applied.

### Heat capacity measurements

The heat capacity of 3 was measured by the relaxation method for temperatures down to 0.3 K and constant magnetic fields from 0 to 7 T, using a Quantum Design PPMS, equipped with a  $^3\text{He}$  cryostat. The experiments were performed on a thin pressed pellet (*ca.* 1 mg) of a polycrystalline sample of 3, thermalized by *ca.* 0.2 mg of Apiezon N grease, whose contribution was subtracted by using a phenomenological expression.

## Results and discussion

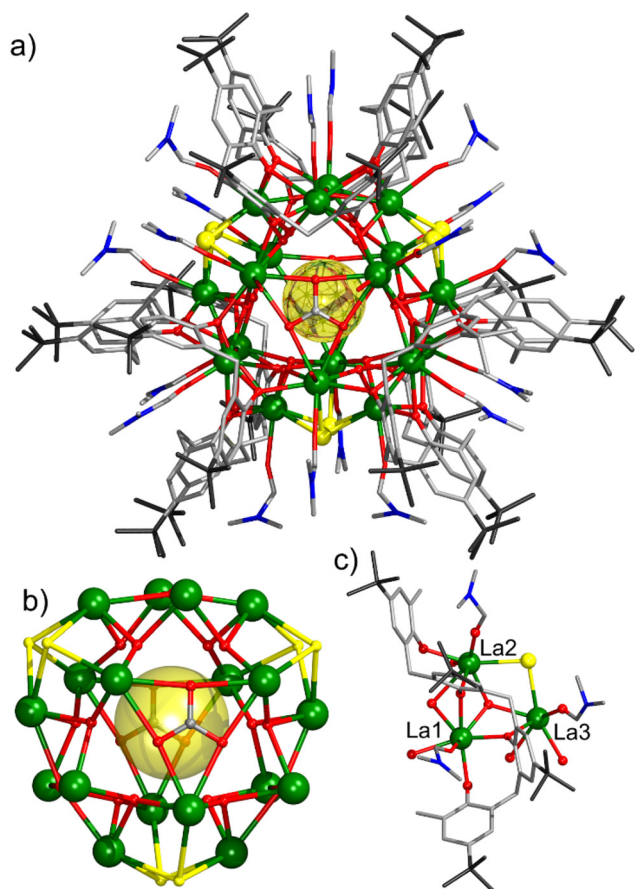
### Structure description

Complexes 1–3 were synthesized by reacting  $\text{Ln}_2^{\text{III}}$ ,  $\text{H}_8\text{TBC}[8]$ ,  $\text{NBu}_4\text{OH}$ , and 1,3,5-tricarboxylic acid in 8 mL of DMF/MeOH/ $\text{CHCl}_3$  ( $v:v:v = 3/3/2$ ) under solvothermal conditions at 120 °C for 48 hours (1 and 2) or at 150 °C for 24 hours (3). We note that the syntheses require the presence of DMF. When conducted under an inert atmosphere, the syntheses produce the same title compounds but in lower yields. The block-like single crystals of  $\text{La}_{18}^{\text{III}}$  (1),  $\text{Nd}_{18}^{\text{III}}$  (2), and  $\text{Gd}_{18}^{\text{III}}$  (3) were isolated by either slow diffusion of hexane into a concentrated solution of the mother liquor containing the product (1, 2), or by slow evaporation of the mother liquor (3), see Experimental section for full details. The crystals were all found to be in the trigonal crystal system and structure solution was performed in the space group  $P\bar{3}1c$  (Tables S1 and S2<sup>†</sup>). The three complexes are structurally analogous, therefore we only provide a representative molecular structural description of 1. The  $\text{La}_{18}^{\text{III}}$  skeleton is fully surrounded by TBC[8] units (Fig. 1). Internally, it is templated by a chloride anion (Fig. 2a) located in the center of the molecule ( $\text{Cl}-\mu_3\text{-OH}$ : 3.47 Å), from which the closest twelve  $\mu_3\text{-OH}^-$  ligands bind ‘internally’ to the eighteen  $\text{La}^{\text{III}}$  ions ( $\text{La}-\text{O}$ : 2.553(7), 2.574(6), and 2.590(6) Å). ‘Externally’  $\mu_3\text{-Cl}^-$  ( $\text{La}-\text{Cl}$ : 2.974(4) Å),  $\mu_4\text{-O}^{2-}$  ( $\text{La}-\text{O}$ : 2.5533(5) and 2.8249(18) Å), and  $\mu_3\text{-CO}_3^{2-}$  ( $\text{La}-\text{O}$ : 2.622(9) Å) ions and the fully deprotonated TBC[8] ligands ( $\text{La}-\text{O}$ : 2.404(9)–2.589(8) Å) connect the metallic core. The TBC[8] ligands are in a twofold double-cone conformation with O-atoms in each phenoxide unit displaying a  $\mu_2$ -binding mode to the La ions (Fig. 1). The remaining coordination sites around the metal center are occupied by terminally bonded  $\text{H}_2\text{O}$  and DMF molecules. The OH counter anion is located at  $\sim 2.6$  Å from the three equivalent coordinated water molecules (Fig. S3a<sup>†</sup>). Bond Valence Sum (BVS) calculations<sup>12</sup> (Table S4<sup>†</sup>) confirm the assignment of  $\text{La}^{\text{III}}$  ions, fully deprotonated TBC[8] ligands, and  $\mu_3\text{-OH}^-/\text{Cl}^-$  ions. The cationic cluster is highly symmetric ( $D_{3h}$ , trigonal planar) with the  $C_3$  axis passing through the two central carbon atoms of the two carbonate ligands (Fig. S3a<sup>†</sup>). La1, La2, and La3 in the asymmetric unit (Fig. 2c) adopt distorted square antiprismatic  $\text{Ln}^{\text{III}}\text{O}_8$ , distorted square antiprismatic  $\text{Ln}^{\text{III}}\text{O}_6\text{Cl}_2$ , and distorted capped square antiprismatic  $\text{Ln}^{\text{III}}\text{O}_8\text{Cl}$  geometries, respectively (Fig. S1<sup>†</sup>).

Hydroxide-bridged lanthanide clusters are typically assembled by the repetition of small building blocks that construct the structural skeleton of the clusters. For example,  $\text{Ln}_{104}^{\text{III}}$  is formed by corner-sharing trigonal bipyramidal

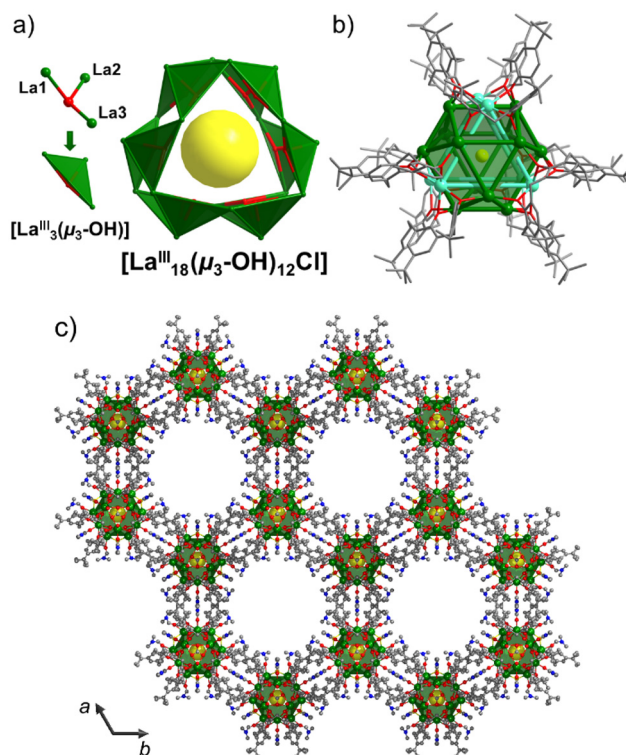






**Fig. 2** (a) Molecular structure of  $[\text{La}_{18}^{\text{III}}(\text{TBC}[8]-8\text{H})_3(\mu_4\text{-O})_3(\mu_3\text{-Cl})_6(\mu_3\text{-OH})_{12}(\mu_3\text{-CO}_3)_2(\text{H}_2\text{O})_6(\text{DMF})_{18}\text{Cl}]^+$ . (b) The  $[\text{La}_{18}^{\text{III}}(\mu_4\text{-O})_3(\mu_3\text{-Cl})_6(\mu_3\text{-OH})_{12}(\mu_3\text{-CO}_3)_2]$  fragment showing the coordination modes of the OH/O/Cl/CO<sub>3</sub> ligands. (c) The asymmetric unit. Color code: Ln: green, O: red, N: blue, Cl: yellow, and C: gray. H atoms, solvent molecules, and OH counter anion are omitted for clarity.

$\{\text{Ln}_5(\mu_3\text{-OH})_6\}$  units and  $\text{Ln}_{60}^{\text{III}}$  by vertex-sharing cubane-like  $\{\text{Ln}_4(\mu_3\text{-OH})_4\}$  moieties. The use of edge- or vertex-sharing triangular  $\{\text{La}_3(\mu_3\text{-OH})\}$  building blocks has been reported in the formation of small clusters with nuclearities ranging from four to twelve, due to the inexistence of templating anions. In this work, the use of  $\text{H}_8\text{TBC}[8]$  and templating chloride ion has enabled us to construct  $\text{Ln}_{18}^{\text{III}}$  *via* assembly of twelve vertex-sharing triangular  $\{\text{La}_3(\mu_3\text{-OH})\}$  motifs that build the metallic core (Fig. 3a). A polyhedral representation of the metallic skeleton (Fig. 3b) describes an intercalated trigonal prism and an irregular polyhedron that consists of six equilateral triangles and eight isosceles trapezia. It is important to note that the release of chloride and carbonate anions during the solvothermal reaction is essential for structuring the metallic skeleton. These anions are slowly produced *in situ* by the dehalogenation of chloroform and the decomposition of DMF, and/or fixation of atmospheric CO<sub>2</sub>. The use of 1,3,5-tricarboxylic acid, although not integrated into the structure, is mandatory. We postulate that it may prevent the isolation of smaller TBC[8]-based Ln complexes, particularly those based on



**Fig. 3** (a) Vertex-sharing of triangular  $\{\text{Ln}_3(\mu_3\text{-OH})\}$  building blocks in the  $[\text{Ln}_{18}^{\text{III}}(\mu_3\text{-OH})_{12}\text{Cl}]$  fragment. (b) Polyhedral representation of the 18-metal-atom core structure as an intercalated trigonal prism (cyan) and an irregular polyhedron (green). (c) Crystal packing of **1** viewed along the *c* axis exhibiting cylindrical channels with a diameter of  $\sim 22$  Å. Color code as in Fig. 2. H atoms, solvent molecules, and OH counter anions are omitted for clarity.

single- or double-deckers. Instead, it facilitates the formation of  $\text{La}_{18}^{\text{III}}$ , the first such structure supported by three TBC[8] ligands.

Analysis of the crystal packing reveals a mesoporous material, whose void space corresponds to  $\sim 50\%$  of its volume (using a  $1.5$  Å spherical probe radius; Fig. S2†).<sup>13</sup> The crystal lattice exhibits cylindrical channels along the *c* axis with a diameter of  $\sim 22$  Å created by the arrangement of  $\text{La}_{18}^{\text{III}}$  molecules in hexagonal motifs mediated by intermolecular  $\text{CCH}_3(\text{tBu})\text{-CCH}_3(\text{tBu})$  contacts of  $\sim 3.9$  Å (Fig. S3b†). The  $\text{OH}^-$  counter anion is H-bonded to the terminally bound water molecules of the cluster ( $\text{La-OH}_2\cdots\text{OH}^-$ :  $2.60$  Å; Fig. S3†).

#### FT-IR, UV-Vis, and TGA-DSC experiments

In the FT-IR spectra of **1-3** (Fig. S4†), the most important bands are related to vibrations associated with  $\nu(\text{C-H})$   $\sim 2953\text{-}2863$   $\text{cm}^{-1}$  ( $\nu(\text{C=O})$   $\sim 1653$   $\text{cm}^{-1}$  (vs) (characteristic of carbonate),  $\nu(\text{arC-C})$   $\sim 1486$   $\text{cm}^{-1}$  (vs) (localized in the TBC[8] ligand; uncoordinated  $\text{H}_8\text{TBC}[8]$  ligand shows two smaller bands at  $1486$  and  $1452$   $\text{cm}^{-1}$ ), and  $\nu(\text{C-O})$  at  $\sim 1296$  (m),  $\sim 1267$  (m), and  $\sim 1206$   $\text{cm}^{-1}$  (s). The absence of the PhO-H vibration band, present in the uncoordinated  $\text{H}_8\text{TBC}[8]$  ligand at  $1246$   $\text{cm}^{-1}$  is another proof of the deprotonation of the



ligand upon coordination. The electronic absorption spectra of 1–3 in CHCl<sub>3</sub> solutions display an intense absorption band at ~304 nm, attributed to  $\pi$ - $\pi^*$  electronic transitions centered on the aromatic rings of the fully deprotonated TBC[8] (Fig. S6†). Their electronic absorption curves are unchanged over a period of 24 hours in CHCl<sub>3</sub> solutions. The TGA curves of 1–3 (Fig. S7–S9†) show a slow release of crystallization solvent up to ~220 °C, wherefrom a rapid loss of mass occurs between 220 and 420 °C, corresponding with the decomposition of the structures. These two mass-loss steps are evidenced by a broad endothermic peak at ~380 °C by differential scanning calorimetry (DSC) in the three complexes. For complexes 1–3, the experimental patterns closely mirror each other, as anticipated for isostructural analogs.

### Magnetic studies

Direct current (dc) magnetic susceptibility and magnetization measurements of 2 and 3 are shown in Fig. 4. At 290 K, the  $\chi_m T$  values of 25.3 (2) and 141.0 cm<sup>3</sup> K mol<sup>-1</sup> (3) are within the anticipated range of 25–28 cm<sup>3</sup> K mol<sup>-1</sup> and 137–142 cm<sup>3</sup> K mol<sup>-1</sup> for eighteen non-interacting Nd<sup>III</sup> and Gd<sup>III</sup> centers, respectively.<sup>14</sup> Upon cooling, the  $\chi_m T$  value of 2 progressively decreases to 9.7 (at 0.1 T) or 9.0 cm<sup>3</sup> K mol<sup>-1</sup> (1.0 T) at 2 K. The  $\chi_m T$  value of complex 3 remains constant between 300 and 50 K, wherefrom it rapidly decreases to 64.9 (0.1 T) or 59.1 cm<sup>3</sup> K mol<sup>-1</sup> (1.0 T) at 2.0 K. The  $\chi_m T$  vs.  $T$  data of 2 is caused by the thermal depopulation of the  $m_j$  energy sublevels of the predominant ground term (<sup>4</sup>I<sub>9/2</sub>) that results from the combined action of electron–electron repulsion and spin–orbit coupling in the Nd<sup>III</sup> center, and the ligand field splitting the  $m_j$  sublevels. At lower temperatures ( $T < 10$  K), the diverging  $\chi_m T$  vs.  $T$  curves at 0.1 and 1 T can be attributed to the Zeeman effect and predominantly antiferromagnetic exchange interactions between the Nd<sup>III</sup> centers. The almost isotropic nature of the

Gd<sup>III</sup> ion ( $s = 7/2$ , <sup>8</sup>S<sub>7/2</sub>) leads to a nearly constant  $\chi_m T$  value at  $T > 100$  K.

Due to the marginal splitting of the four doublet energy states of the <sup>8</sup>S<sub>7/2</sub> ground term, the distinct drop of  $\chi_m T$  at  $T < 50$  K can only be attributed to the predominantly antiferromagnetic exchange interactions between the Gd<sup>III</sup> centers. Fitting the  $\chi_m$  data of 3 to the Curie–Weiss law yields the Weiss temperature  $\theta = -3.1$  K, denoting weak yet significant antiferromagnetic correlations, as also confirmed by heat capacity ( $c_p$ ) measurements (Fig. S11†). The zero-field magnetic contribution to  $c_p$  can be well described by the Schottky model for eighteen spins  $s = 7/2$  per cluster and an effective magnetic field  $B_{\text{eff}} = 0.6$  T that mimics the magnetic interactions (notably,  $g\mu_B B_{\text{eff}} = 2.8$  K  $\approx |\theta|$ , for  $g = 2.0$ ). Considering the isomorphous nature of 2 and 3, it is highly probable that antiferromagnetic interactions also exist in 2. At 2.0 K, the molar magnetization  $M_m$  vs.  $B$  plots of 2 and 3 show an approximately positive linear dependence and reach 20.2 (2) and 125.4  $N_A \mu_B$  (3), respectively, at 5.0 T (roughly 34% of the saturation value 58.9  $N_A \mu_B$  for 2). This is due to the measurement of powdered samples, *i.e.*, the determination of the mean value of randomly oriented crystallites, and the additional contribution of magnetically anisotropic Nd<sup>III</sup> centers in complex 2. The magnetization data further confirm the presence of predominantly antiferromagnetic exchange interactions in both compounds.

The MCE of 3 was investigated by recording the isothermals (2–13 K) of the magnetization data in an applied field (Fig. S10†) and the heat capacity at a constant field between 0.3 and 30 K (Fig. S11†). Calculation of the temperature and field dependences of the magnetic entropy change from  $M_m$  data uses the Maxwell equation:

$$\Delta S_m(T, \Delta B) = \int_0^{\Delta B} \left( \frac{\partial M_m}{\partial T} \right)_B dB',$$

for any applied magnetic field change  $\Delta B = B - 0$ .<sup>5,15,16</sup> The  $c_p$  data provide better information in terms of the MCE since they allow us to determine both  $\Delta S_m(T, \Delta B)$  and  $\Delta T_{\text{ad}}(T, \Delta B)$  changes, straightforwardly from the entropy (Fig. S10†), calculated as:

$$S(T, B) = \int_0^T \frac{c_p}{T'} dT'.$$

The corresponding changes of  $\Delta S_m$  and  $\Delta T_{\text{ad}}$  at different  $T$  and  $\Delta B$  are shown in Fig. 5. Both approaches, based on calculating  $\Delta S_m$  from  $M_m$  and  $c_p$  data, respectively, provide identical results. We find that both  $-\Delta S_m$  and  $\Delta T_{\text{ad}}$  improve gradually by increasing  $\Delta B$ , reaching the values of 23.7 J K<sup>-1</sup> kg<sup>-1</sup> and 8.6 K, respectively, at 2.2 K and  $\Delta B = 7$  T. The maximum experimental entropy change corresponds to *ca.* 77% of the available entropy content, *i.e.*,  $18 \times R_s \ln(2s + 1) = 30.8$  J K<sup>-1</sup> kg<sup>-1</sup> ( $R_s = R/M$ ). The MCE is surprisingly high considering the supramolecular nature of the TBC[8] ligand – with its low magnetic density given by  $M_w/N_{\text{Gd}}$ , where  $M_w$  is the molecular weight and  $N_{\text{Gd}}$  is the number of gadolinium metal ions constituting

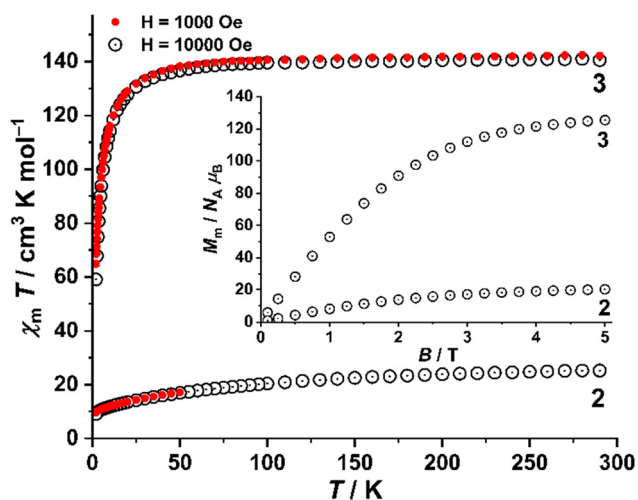


Fig. 4 Susceptibility data for 2 and 3 in static magnetic fields:  $\chi_m T$  vs.  $T$  of 2 and 3 at 0.1 (filled red circles) and 1.0 T (open black circles). Inset: molar magnetization  $M_m$  vs. applied magnetic field  $B$  at 2.0 K.



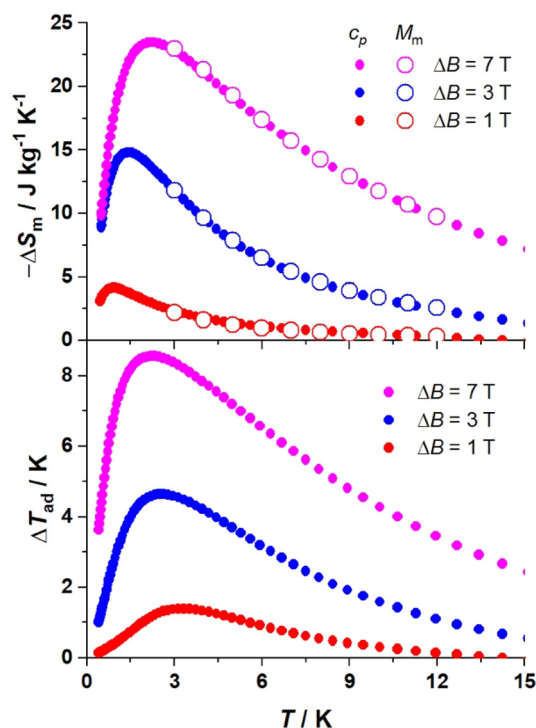


Fig. 5 Magnetic entropy change (top) and adiabatic temperature change (bottom) for **3**, obtained from the molar magnetization (open symbols) and heat capacity (filled symbols) data, shown in Fig. S10 and S11,<sup>†</sup> respectively, for selected applied field changes, as labeled.

the cluster. Although the magnetocaloric response of **3** is not in close range to the best molecular magnetorefrigerants based on polynuclear homometallic lanthanide clusters,<sup>1a,b,2a,16</sup> it represents the best result obtained with calix[*n*]arene-based molecular clusters.<sup>17</sup>

## Conclusions

In summary, the use of the “anion template strategy” and the involvement of “templating ligands” – hitherto neglected in calix[*n*]arene coordination chemistry – allowed us to synthesize a family of highly symmetric ( $D_{3h}$ , trigonal planar)  $\text{Ln}_{18}^{\text{III}}$  coordination clusters. The title compounds exhibit structural features common to high-nuclearity homometallic lanthanide hydroxy clusters obtained with smaller ligands (*e.g.*,  $\text{Gd}_{140}^{\text{III}}$ ,  $\text{Ln}_{104}^{\text{III}}$ ,  $\text{Dy}_{76}^{\text{III}}$ ) but uncommon in calix[*n*]arene-based clusters, such as high symmetry, the presence of templating anions, and the presence of an inner core constructed by the repetition of small  $\{\text{Ln}_x(\mu_n\text{-OH})_y\}$  building blocks. The crystal packing of the molecules creates a material with cylindrical channels of  $\sim 22$  Å diameter along the *c* axis of the unit cell. Based on the size of the solvent channels, the title compounds could be of interest for the development of materials for sorption experiments. Investigation of the magnetocaloric effect on the gadolinium derivative yields a  $-\Delta S_m = 23.7 \text{ J K}^{-1} \text{ kg}^{-1}$ ; this value shows the potential of engineering magnetic refrigerants with calix[*n*]arene ligands. Current

work focuses on (a) the synthesis of  $\text{Gd}_{18}^{\text{III}}$  analogs with C[8] (*i.e.*, exchanging ‘Bu groups in the ligand to hydrogens) to increase the magnetic density ( $M_w/N_{\text{Gd}}$ ) and thus the magnetic entropy content, and (b) the application of the reported strategies targeting larger homometallic lanthanide clusters. Results stemming from the ongoing work will be communicated in due course and may lead to a better understanding of the mechanism involved in the construction of high-nuclearity lanthanide clusters with TBC[8] and the relationship between structural modifications and magnetic entropy changes.

## Conflicts of interest

There are no conflicts to declare.

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