

Polypyrrole-derived Activated Carbons for High-performance Electrical Double-layer Capacitors with Ionic Liquid Electrolyte

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As electrical energy storage and delivery devices, carbon-based electrical double-layer capacitors (EDLCs) have attracted much attention for advancing the energy-efficient economy. Conventional methods for activated carbon (AC) synthesis offer limited control of their surface area and porosity, which results in a typical specific capacitance of 70–120 F g⁻¹ in commercial EDLCs based on organic electrolytes and ionic liquids (ILs). Additionally, typical ACs produced from natural precursors suffer from the significant variation of their properties, which is detrimental for EDLC use in automotive applications. A novel method for AC synthesis for EDLCs is proposed. This method is based on direct activation of synthetic polymers. The proposed procedure allowed us to produce ACs with ultrahigh specific surface area of up to 3432 m² g⁻¹ and volume of 0.5–4 nm pores up to 2.39 cm³ g⁻¹. The application of the produced carbons in EDLCs based on IL electrolyte showed specific capacitance approaching 300 F g⁻¹, which is unprecedented for carbon materials, and 5–8% performance improvement after 10 000 charge–discharge cycles at the very high current density of 10 A g⁻¹. The remarkable characteristics of the produced materials and the capability of the fabricated EDLCs to operate safely in a wide electrochemical window at elevated temperatures, suggest that the proposed synthesis route offers excellent potential for large-scale material production for EDLC use in electric vehicles and industrial applications.

1. INTRODUCTION

Electrical double-layer capacitors (EDLCs) are a unique type of high-power electrochemical energy storage devices, where the capacitance arises from the charge separation at an electrode-electrolyte interface.^[1] Recently, EDLCs have attracted considerable attention because of their high power density, sub-second charging and ultra-long cycle life, which make them attractive for a wide range of applications, including consumer electronics, uninterruptable power sources, energy efficient industrial equipment, power grid systems and most importantly hybrid electric vehicles.^[2–8]

The energy storage in EDLC is based on the adsorption of electrolyte ions on the large specific surface area of electrically conductive porous electrodes, most commonly porous carbons. Several promising methods have been explored for porous carbon synthesis for industrial EDLC applications. For example, the zeolite-templated carbons studied by Kyotani^[9] and later by other groups^[10–13] allow formation of uniform microporous carbons with high specific capacitance of 90–150 F g⁻¹ in organic electrolytes. However, the current need to use dangerous hydrofluoric (HF) acid for efficient template dissolution, has so far prevented commercial applications of this technology. Selective etching of metals from metal carbides offer controlled pore size distribution in the produced carbide-derived carbons (CDCs).^[14–16] Specific capacitance of 70–170

F g⁻¹ was demonstrated in organic electrolytes and ionic liquids,^[5,17–19] but the material commonly suffers from tortuous pore shape, unless CDCs are produced with the help of ordered silica templates.^[20,21] Exfoliation of graphite with subsequent activation is yet another interesting approach for porous carbon formation, which was recently demonstrated to offer impressive capacitance of up to 167 F g⁻¹ in ionic liquid/organic solvent mixture.^[7] However, poor control over the pore size distribution (PSD) and formation of significant content of large mesopores in this carbon limit its volumetric capacitance. Furthermore, as shown by prior studies^[5,22] the specific capacitance in small microporous pores, where the ion solvation shell becomes highly distorted,^[5,23,24] could be larger than in mesopores. Strictly mesoporous carbons, such as carbon onions and carbon nanotubes^[3,25] may find applications when ultra-fast rate capability is needed. However, their limited surface area and excessive porosity minimizes the energy density of the fabricated devices.

High surface area activated carbons (ACs) are predominant electrode materials for commercial EDLCs due to their well developed manufacturing technologies, easy production of large quantities, relatively low cost and great cycle stability. While commercial ACs exhibit only moderately high specific capacitance (commonly 70–120 F g⁻¹ in organic electrolytes)^[2,26–28] the recent developments in their synthesis^{29–33} demonstrate that for the significant portion of the EDLC applications

ACs will likely remain the material of choice, in spite of the alternative structures introduced^{2-13,18-21,25}. Therefore, novel methods of AC synthesis with well controlled properties, high volume of micropores and better capacitive characteristics are of the highest demand for further development of EDLC technology. In our recent paper we utilized an environmentally friendly low-temperature hydrothermal carbonization³⁴⁻³⁷ followed by activation to transform natural organic materials into ACs with well developed porosity (surfaced area of 2100-2450 m² g⁻¹) and very high capacitance (140-210 F g⁻¹) in organic electrolytes³². The serious disadvantage of natural organic precursors (such as wood, seaweed, coconut shells and others), however, is the lack of sufficient uniformity and consistency in their properties, which could be prohibitive for automotive applications of EDLCs. Furthermore, the use of flammable organic solvents in EDLCs raises many safety concerns. In contrast, ionic liquids (ILs) electrolytes offer greatly enhanced operation safety, higher voltage stability (and thus higher energy density) and longer lifetime to EDLCs^{18,38-41}.

Here, we report for the first time on the preparation of ACs from synthetic polymers by a one-step chemical activation process with KOH. For the proof-of-concept studies we selected polypyrrole (PPy) as synthetic precursor material due to a high percentage of carbonaceous residue left after thermal treatment. We demonstrate that PPy could be transformed to highly porous carbons with BET^[42] specific surface area (SSA) of up to 3432 m² g⁻¹, one of the highest values ever reported for carbon materials.^[7] We further demonstrate that PPy-derived AC offer truly outstanding specific capacitance approaching ≈ 300 F g⁻¹ in 1-ethyl-3-methylimidazolium tetrafluoroborate (EMImBF₄) electrolyte at 60 °C. These values are 100 to 300% higher than specific capacitance reported in commercial ACs, CDC, carbon nanotubes (CNT), and graphene, thus representing a significant technological breakthrough enabled by the proposed chemical synthesis route.

2. RESULTS AND DISCUSSION

The carbon formation procedure employed in our study involved the room temperature synthesis of the polymer polypyrrole from an aqueous suspension of pyrrole and ferric chloride solution. The resulting polypyrrole had a carbon content (wt%) of 65 with 16.6, 13.5 and 4.5 of oxygen, nitrogen and hydrogen respectively (Table S1, Supporting Information). The AC preparation procedure performed at 600–700 °C utilizes these uniformly distributed hetero atoms (primarily O and N) to transform the polypyrrole into highly porous carbons with tunable elemental composition and interconnected pore network. The produced carbons were crushed into finer powders and labeled according to their activation temperature as AC-P600, AC-P650, and AC-P700.

The produced ACs are made up of irregularly shaped platelets 20-50 μm in lateral dimensions and 5-10 μm in thickness, with sharp corners and “swiss cheese”-like macropores of 5 μm in diameter, originating from the gas bubble formation during the activation (Figure 1a). The activation changed the original aerogel-like structure of the polypyrrole precursor (Figure S1, Supporting Information) completely.

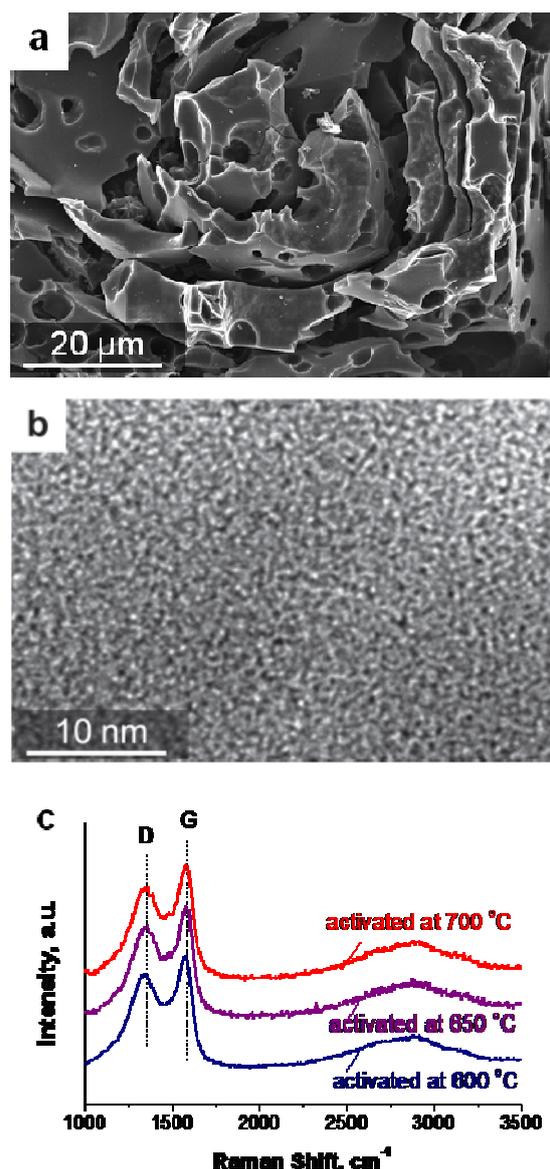


Figure 1. Structure of polypyrrole-derived activated materials: a) Scanning electron microscopy (SEM) and b) TEM images of the carbon activated at 700 °C and c) Raman spectra of samples activated at different temperatures.

High-resolution transmission electron microscopy (TEM) revealed highly disordered porous carbon structure without graphite ribbons and crystalline impurities (Figure 1b). Raman spectroscopy studies similarly showed spectra characteristic for highly disordered carbons^[10,20,21,43] (Figure 1c). In the spectral range from 0 to 2000 cm⁻¹ a perfect graphite shows only one narrow Raman active mode, the G-band, at 1582 cm⁻¹ and corresponding to graphite in-plane vibrations.^[44-46] Disordered carbons additionally show the D band at ~ 1350 cm⁻¹. This band is associated with a double-resonance Raman process in disordered carbon,^[46,47] when any phonon mode irrespective of the symmetry can give rise to this peak if the wave vector of the phonon is twice as large as the wave vector of the electronic transition excited by the incident phonon. The large value of the full width at half maximum (FWHM) of D and G bands (~ 190 and 90 cm⁻¹, respectively, for all samples) and a

high value of the ratio of integrated intensities of these peaks ($I_D/I_G=1.8$) is typical for porous carbons with very high degree of structural (i.e., graphitic) disorder.^[10,20,21,43] The lack of the sharp and strong peaks in the spectral range 2400 - 3300 cm^{-1} (2D, D+G, T+D, T+G, 2D' and other bands)^[46] similarly suggest the lack of graphitic ribbon structures in the sample,^[46] confirming the XRD measurements (Figure 1b).

The nitrogen sorption isotherms of the ACs (Figure 2a) showed significant impact of the activation temperature on their porosity. The AC-P600 sample activated at the lowest temperature of 600 °C exhibits a type-I sorption isotherm (in the Brunauer classification)^[42] with saturation at a relative pressure (P/P_0) of ca. 0.2, characteristic for microporous materials with very low volume of pores >2 nm. The total amount of N_2 adsorbed at $P/P_0 \sim 0.99$ is $\sim 670 \text{ cm}^3 \text{ g}^{-1}$, which corresponds to the moderately high pore volume of $1.1 \text{ cm}^3 \text{ g}^{-1}$. As the activation temperature increases to 650 and 700 °C the pore volume doubles and the isotherms change to a type-IV with a significant slope at higher relative pressures, commonly associated with capillary condensation in small mesopores. The total BET (Brunauer Emmett Teller)^[42] SSA (specific surface area) of the carbons increases with activation temperature from 2095 to 3246 and to $3432 \text{ m}^2 \text{ g}^{-1}$. These surface areas are higher than those of porous carbons produced by activation of graphene,^[7] higher than theoretical surface area of graphene ($2630 \text{ m}^2 \text{ g}^{-1}$) and over 7 times higher than that of porous carbons hydrothermally synthesized with the assistance of sacrificial templates reported previously.^[48] This high surface areas likely result from the formation of multiple defects (holes) in the individual graphene layers of the produced carbons, which lead to a larger number of surface sites available for N_2 adsorption. According to non-local density functional theory (NLDFT) calculations, sample AC-P600 contained virtually no pores larger than 3 nm (Figure 2b, c). However, activation at higher temperatures of 650 and 700 °C leads to the noticeable increase in the volume of small mesopores in the range of 2-4 nm. Table 1 summarizes the porosity characteristics of our samples. The difference in absolute values between BET and NLDFT SSA (Table 1) is commonly observed and is expected as the calculations are based on different assumptions.^[20,42]

Following Kurig et al.^[40] and our prior work^[20,41] we selected EMImBF₄ as a very promising IL electrolyte. Specific capacitance of the produced carbon samples was estimated using cyclic voltammetry (C-V) at different scan rates (Figure 3). At low scan rate of 1 mV s^{-1} , the C-V curves show quasi-rectangular shape even at room temperature, which is characteristic for an ideal EDLC with little electrolyte diffusion limitations. The pseudocapacitance peaks visible at $\sim 1.2 \text{ V}$ in the symmetrical EDLC may result either from impurities present in the IL or from the reaction of the IL with the functional groups on the carbon surface. Importantly for practical applications, the Faradaic reactions present did not negatively affect the cycle stability of the assembled EDLC. As the sweep rate increased to 100 mV s^{-1} the pseudocapacitance peaks disappeared (Figure 3 a, b) but the shape of the C-V curves retained nearly rectangular shape, particularly when measured at 60 °C on samples activated at 650 and 700 °C (Figure 3b). Within the 20-60 °C measurement temperature range samples AC-P650 and AC-P700 retained up to 60% of the maximum specific capacitance when the sweep rate was increased from 1 to 100 mV s^{-1} (Figure 3c, d). The decrease in capacitance at higher rates is common for activated carbons and originates

from the insufficient time available for ion diffusion and adsorption inside the smallest inner pores within the carbon particles. Considering the large ($\sim 300 \mu\text{m}$) electrode thickness and very large (un-optimized) particle size (Figure 1a) the 60% capacitance retention demonstrated a very good promise of the synthesized materials for high-power applications. Note that particle size reduction to sub-micron range¹⁹ should result in the sufficient (10-fold or more) decrease in the diffusion length ℓ_d for ions accessing the inner core of the carbon particles surface through the network of 0.5-4 nm pores (Figure 2 b, c), which, in turn, should significantly reduce the characteristic charging time ($\tau \propto \ell_d^2$) and dramatically improve the EDLC power characteristics.

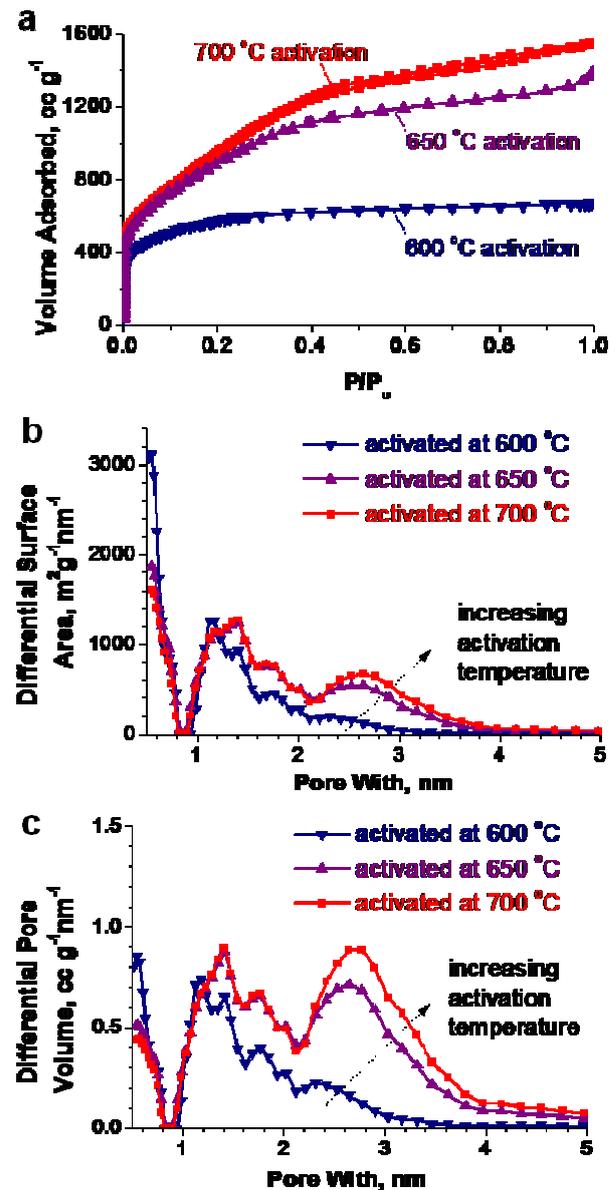


Figure 2. Porosity characterization: a) N_2 gas adsorption isotherms and b,c) NL-DFT pore size distribution of polypyrrole-derived ACs.

Table 1. Porosity properties of the polypyrrole-derived activated carbon materials

| Sample | BET-SSA ^[a] (m ² g ⁻¹) | DFT-SSA ^[b] (m ² g ⁻¹) | MP-SSA ^[c] (m ² g ⁻¹) | MPV ^[d] (cm ³ g ⁻¹) | TPV ^[e] (cm ³ g ⁻¹) |
|-------------------------------|---|---|--|--|--|
| AC-P600 (Activated at 600 °C) | 2095 | 1703 | 1550 | 0.71 | 1.10 |
| AC-P650 (Activated at 650 °C) | 3246 | 2160 | 1520 | 0.84 | 2.13 |
| AC-P700 (Activated at 700 °C) | 3432 | 2244 | 1429 | 0.81 | 2.39 |

[a] BET Specific Surface Area measured for $P/P_0=0.06-0.14$, [b] NL DFT Specific Surface Area, [c] NL DFT Specific Surface Area of Micropores (pores <2 nm), [d] NL-DFT Micropore Volume, [e] Total Pore Volume measured at $P/P_0=0.994$.

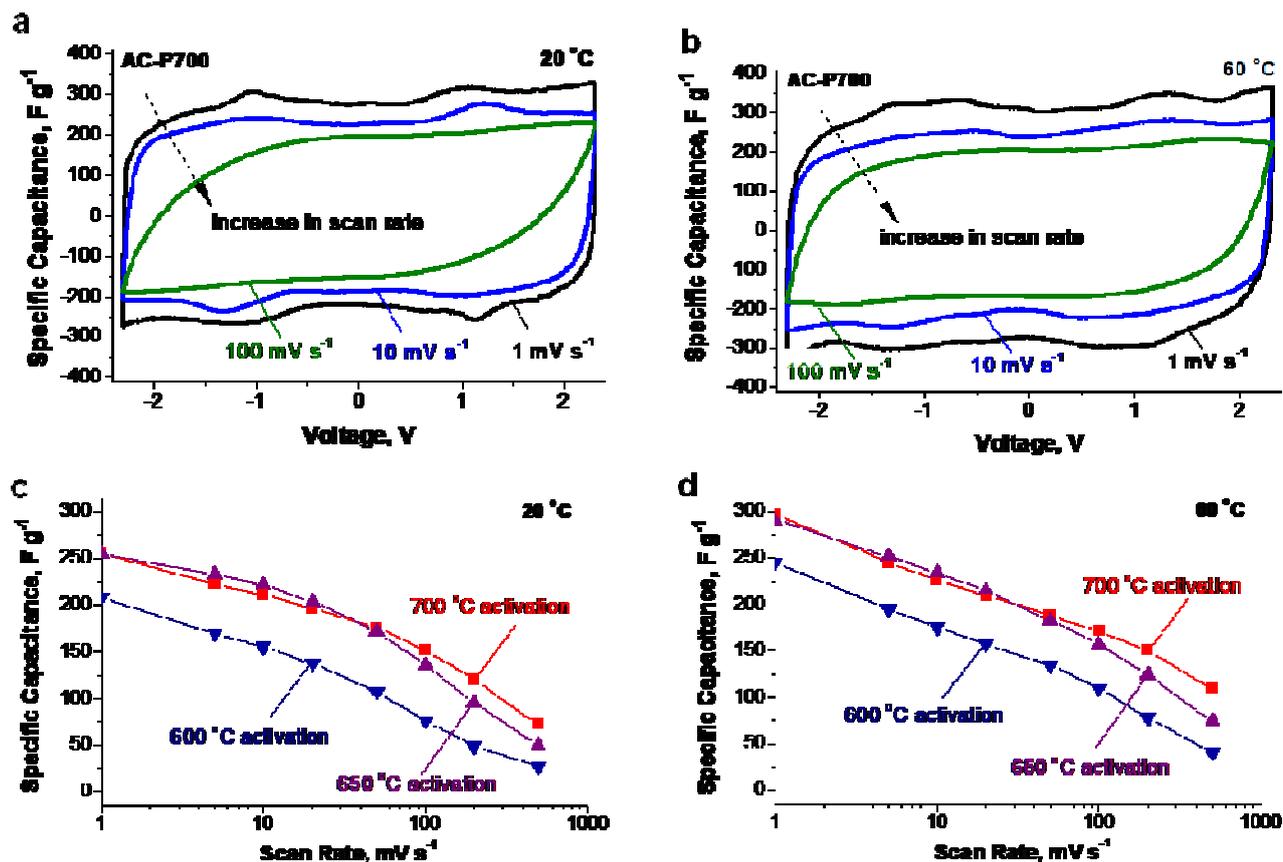


Figure 3. Electrochemical characterization of polypyrrole-derived AC materials in EMImBF₄ (1-ethyl-3-methylimidazolium tetrafluoroborate) ionic liquid at room temperature (20 °C) and at 60 °C: a, b) cyclic voltammograms of carbon activated at 700 °C and c, d) specific capacitance of the carbons activated at three different temperatures as a function of scan rates.

The most impressive result of our studies was the discovery of the enormously high specific capacitance of the present activated carbons. The obtained capacitance values (up to 256 F g⁻¹ at 20 °C and above 300 F g⁻¹ at 60 °C) are unmatched by the current technology and exceed by 100-200 % the capacitance of conventionally produced activated carbons^[41] and by over 50% the capacitance of templated CDC with a total pore volume and pore size distribution similar to that of our samples^[20] (all obtained with the same electrolyte). The very high SSA of the produced samples (Table 1) is unlikely to be the single origin of their outstanding performance. In fact, the SSA-normalized capacitances for our samples calculated at the 1 mV s⁻¹ rate were in the range of 7.6-13.2 μF cm⁻² (for AC-P700) to 10-14.4 μF cm⁻² (for AC-P600) (Figure S2, Supporting Information, SI), which is noticeably higher than the 7.3

μF cm⁻² previously observed with templated mesoporous CDC^[20]. Because both the pore size distribution and the volume of sub-nm pores in our samples is comparable to that of mesoporous CDC^[20] and the observed pseudocapacitive peaks are relatively small (Figure 3a,b), the noticeably higher specific capacitance of polypyrrole-derived activated carbons may emphasize the importance of the carbon defects-electrolyte interactions for the capacitance enhancement, as previously observed in both organic^[21,25] and aqueous^[49] electrolytes. By studying strictly microporous CDC in ILs,^[18] Largeot et al. observed that decreasing the CDC synthesis temperature from 1000 to 550 °C and simultaneously decreasing the average pore size from 1.1 to 0.7 nm resulted in the significant increase of the SSA-normalized capacitance from ~7 to 13.4 μF cm⁻². The authors suggested that the optimum pore size should

mirror the ion size for the most efficient adsorption¹⁸ and that larger than optimum pores should reduce the gravimetric capacitance.^[18] They ignored, however, the effect of synthesis temperature on the microstructural changes in carbon. Our studies clearly demonstrate that SSA-normalized capacitance in excess of $14.4 \mu\text{F cm}^{-2}$ can be achieved in largely mesoporous materials with the average pore size of $\sim 2 \text{ nm}$ (Figure 1). Therefore, the presence of high volume of sub-nm pores, which may slow down the ion diffusion rate,^[20] is not needed for high energy storage characteristics of EDLC based on ILs.

The higher capacitance achieved at higher testing temperatures (compare Figure 3c and 3d) is likely related to the reduction in the viscosity of EMImBF₄ from 34 cP at room temperature to 12.7 cP at 60 °C and the increase in its ionic conductivity from 12.0 mS cm^{-1} at room temperature to 39.6 mS cm^{-1} at 60 °C.

The Nyquist plots of AC-P700 and AC-P650 (Figure 4 a, b) exhibit the typical features of porous electrodes with a relatively short 45° Warburg region at high-medium frequencies, and an almost vertical line at low frequencies, where the behaviors become mainly capacitive. The 45° segment in the Nyquist plot is related to the diffusion of the ions into the bulk of electrode particles. The decreased length of these segments was observed when the temperature was increased from 20 to 60 °C (compare Figure 5a and b), indicating reduced resistance encountered by the ions during their transport into the particle core at higher temperatures. Sample AC-P600 showed the largest deviation from the ideal behavior and the largest ionic resistance (Figure 5 a, b), suggesting that an activation temperature of 600 °C is not sufficient to achieve the well-developed porosity required for fast IL ion diffusion.

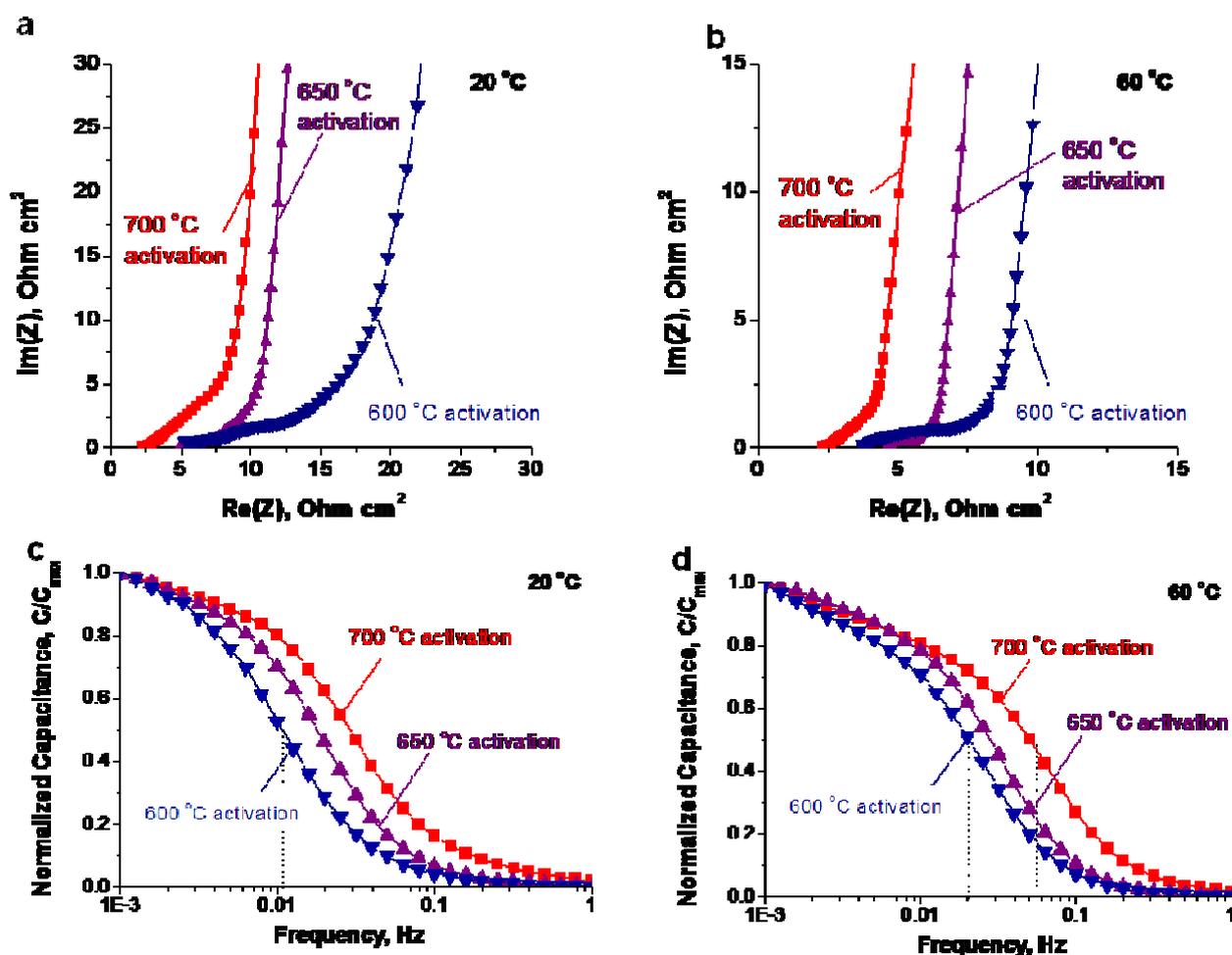


Figure 4. Electrochemical characterization of polypyrrole-derived activated carbon materials activated at various temperatures measured in EMImBF₄ ionic liquid at room temperature (20 °C) and at 60 °C : (a, b) Nyquist plots, (c, d) frequency response.

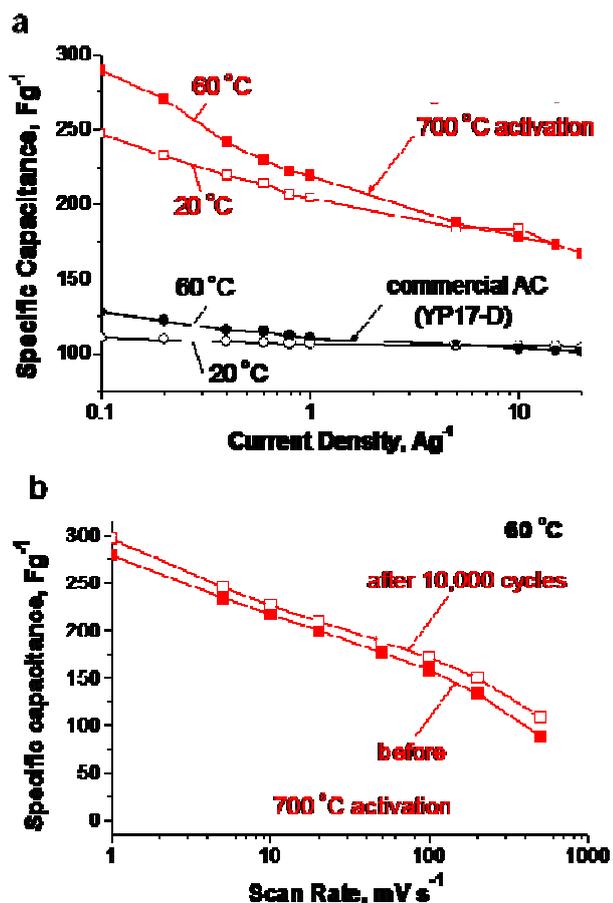


Figure 5. Electrochemical characterization of polypyrrole-derived activated carbon activated at 700 °C in EMImBF₄ ionic liquid: (a) capacitance retention with current density at room temperature (20 °C) and at 60 °C compared to that of commercially available activated carbon YP17-D, (b) capacitance retention/improvement after 10,000 galvanostatic charge-discharge cycles.

The performance of EDLCs in real applications is primarily determined by their charge-discharge (C-D) characteristics, which reveal their energy and power performance. Figure 5a summarizes the effect of the increasing current loads (from 100 mA g⁻¹ to 20 A g⁻¹) in galvanostatic charge-discharge tests on the specific capacitance of the most promising carbon electrode sample, AC-P700. An ideal EDLC with an infinitely fast ion transport should deliver the same energy at virtually any current density. The experimentally observed decrease of capacitance at higher current (most visible in a logarithmic scale, Figure 5a) is generally attributed to the resistance of ions traveling within the nanopores and interacting with carbon functional groups and defects. In spite of the large particle size, AC-P700 electrode showed only a moderate decrease in capacitance from 290 to 167 F g⁻¹ at 60 °C when the current density was increased to ultra-high value of 20 A g⁻¹. While we expect to achieve significantly better performance with particle size and electrode thickness reduction, the obtained results are already noticeably superior than that of recently reported high-capacitance mesoporous carbons prepared using more traditional synthesis routes. For example, the mesoporous carbons produced by conventional pyrolysis and activation of citrates^[50] showed an impressive specific capacitance approaching 180 F g⁻¹ in EMImBF₄ at 60 °C at the low-

est current density. However, this capacitance dropped to below 50 F g⁻¹ as the current density increased to 10 A g⁻¹. In another interesting study on EDLCs employing IL electrolytes,^[51] mesoporous AC fibers were found to exhibit a specific capacitance approaching 196 F g⁻¹ at an ultra-low current density, but only 100 F g⁻¹ at a slightly higher but still very moderate current density of 0.5 A g⁻¹. Even templated CDC^[20] with sub-mm particle size and straight, aligned mesopores demonstrated a reduction of its capacitance from their maximum value of 140-170 F g⁻¹ to only 20-60 F g⁻¹ when the current density approached 20 A g⁻¹. A commercial activated carbon, YP-17D, optimized for use in high-power supercapacitor applications exhibited promising capacitance retention, but suffered from significantly lower capacitance values (Figure 5a). To the best of our knowledge the achieved carbon capacitance values at both small (0.1 A g⁻¹) and large (20 A g⁻¹) current densities are unprecedented for IL electrolytes.

We would like to emphasize that in spite of the redox reactions evidently present (Figure 3a, b), all the assembled EDLC devices demonstrated outstanding cycle stability. In fact, after 10,000 charge-discharge galvanostatic cycles between 0 and 2.3 V performed at 60 °C at a very high current density of 10 A g⁻¹, the specific capacitances of the electrodes commonly increased by 5-8% of their original values (Figure 5b), indicating excellent cycle stability of the polypyrrole-derived activated carbon AC-EMImBF₄ electrochemical system. Moreover, when the EDLC devices were tested in higher potential ranges (2.3–3 V), the CV curves still preserved rectangular shape. The typical C-D profiles remained linear (Figure S4b, Supporting Information) and the IR drop (linked to the ESR) at a high current density of 1 A g⁻¹ remained moderately small. These experiments demonstrate the PPy-based AC electrodes in an EMImBF₄ electrolyte to possess stable capacitive performance in a wide electrochemical window and to offer safe and stable operation at 60 °C.

3. CONCLUSIONS

In conclusion, we have reported a novel promising approach to produce porous carbons from synthetic polymer precursors for use in advanced EDLC systems. Synthesis of polypyrrole at room temperature followed by activation at 600–700 °C allowed us to produce porous carbons exhibiting pore volume in the range of 1.1–2.39 cm³ g⁻¹ and BET surface area in the range of 2095–3432 m² g⁻¹. Activation at higher temperatures lead to increase in the surface area of the carbons and the broadening of the pore size distribution through the formation of additional volume of 1–4 nm pores. When tested in a symmetric EDLC configuration with an EMImBF₄ ionic liquid electrolyte, the synthesized carbons demonstrated specific capacitance of up to 300 F g⁻¹, which is more than a two-fold improvement compared to state-of-the-art commercial carbons and other recently reported novel carbon materials. Charge-discharge tests on the assembled EDLCs performed at 60 °C showed no visible degradation and even small (up to 8%) increase in the specific capacitance after 10 000 galvanostatic cycles. Very good rate performance at high scan rates and current densities suggested a great promise of the proposed technology for safe and high-energy EDLC applications.

4. EXPERIMENTAL SECTION

Activated Carbon Preparation. Polypyrrole (PPy) is a conducting polymer formed from a number of connected pyrrole ring structures. The polypyrrole was prepared as follows; 3 g of pyrrole (recently distilled under nitrogen) was added to 200 mL of 0.5 M FeCl₃ solution and the mixture was magnetically stirred for 2 h. The generated polypyrrole was recovered by filtration, thoroughly washed with distilled water and dried. The dry polypyrrole was then activated by heating (at a ramp rate of 3 °C/min) a KOH/PPy mixture (at KOH/PPy weight ratio of 4) under nitrogen to a final temperature of 600, 650 or 700 °C for 1 h. The activated carbons were then thoroughly washed with 10 wt% HCl to remove any inorganic salts, and then with distilled water until neutral pH and finally dried in an oven (120 °C).

Carbon Characterization. The morphology of the prepared carbon materials was observed via a Leo 1530 (LEO, Osaka, Japan, now Nano Technology Systems Division of Carl Zeiss SMT, USA) scanning electron microscope (SEM). TEM observations of the AC microstructure were carried out using a Philips CM200UT microscope (Philips, Netherlands). Nitrogen sorption isotherms of the carbon samples were collected at -196 °C using a Micromeritics ASAP 2020 sorptometer. The surface area was calculated using the BET method based on adsorption data in the relative pressure (P/P_0) range 0.06 to 0.14 and total pore volume was determined from the amount of nitrogen adsorbed at a relative pressure (P/P_0) of 0.99. The pore size distribution (PSD) was determined via a Non Local Density Functional Theory (NLDFT) method using nitrogen adsorption data, and assuming a slit pore model. The Raman spectra were recorded using a Nicolet Almega XR Micro and Macro Raman Analysis System (Thermo Fisher Scientific Inc., USA). The source of radiation was laser operating at a wavelength of 488 nm and sample exposure time of 5 mins.

Device Assembly. Electrodes were prepared by mixing 94 wt% carbon samples and 6 wt% polytetrafluoroethylene (PTFE) in ethanol to form a slurry. The slurry was stirred uniformly and heated at the same time on a hot plate at 150 °C to evaporate most of the ethanol and form a plasticine-like material. The resulting AC-PTFE composite was then rolled to ~300 μm thick electrode. The electrode was placed into a vacuum oven (100 °C) overnight to remove moisture and residual hydrocarbons, and then was cut into a circular shape with diameter of 1/2 inch for EDLC electrodes. 2016 stainless-steel coin cells with two symmetrical carbon electrodes separated by two GORE™ PTFE separators (W.L Gore and Associates, USA) of ~25 μm in thickness and ~60% porosity were assembled inside an Ar-filled glove box (<1 ppm of oxygen and H₂O, Innovation Technology, USA). Al foil of 300 μm in thickness roughened using a 600 grit SiC sandpaper and coated by a thin layer (~10 to 20 μm) of carbon coating (BW 525, Superior Graphite, USA) was attached to each electrode and served as current collector. Carbon coating was used to reduce the interfacial resistance between the electrode and the current collector. Ionic liquid EMImBF₄ (purity >98%, Ionic Liquids Technology, Germany) was selected as electrolyte because of its high dielectrical permittivity combined with relatively high ionic conductivity and low viscosity at <60 °C.

Electrochemical Testing. Cyclic voltammetry (CV) studies were performed using a Solartron 1480A MultiStat (Solartron Analytical, UK) in the voltage range -2.3 V to +2.3 V and in

scan rates from 1 to 1000 mV s⁻¹. The gravimetric capacitance, C (F g⁻¹), was calculated according to

$$C = \frac{2I}{(dV/dt)m}$$

where I is the current (A), dV/dt is the scan rate (V s⁻¹), and m is the mass (grams) of carbon in each electrode.

Electrochemical impedance spectroscopy (EIS) measurements were carried out using a Gamry Reference 600™ Potentiostat/Galvanostat/ZRA (Gamry Instruments, Inc., USA) in the frequency range of 1 mHz-100 kHz with a 10 mV AC amplitude. The gravimetric capacitance, C (F g⁻¹), was calculated according to

$$C = \frac{2 \cdot |\text{Im}(Z)|}{2\pi f \cdot [(\text{Im}(Z))^2 + (\text{Re}(Z))^2]m}$$

where f is the operating frequency (Hz), $\text{Im}(Z)$ and $\text{Re}(Z)$ are the imaginary and real parts of the total device resistance (Ohm), and m is the mass (grams) of carbon in each electrode.

Galvanostatic charge-discharge cycle tests (GC) were measured using an Arbin BT-2000 testing system (Arbin Instruments, USA) in the voltage range 0-2.3 V and at charge-discharge current densities between 0.1 and 20 A g⁻¹, based on the mass of a single electrode. The gravimetric capacitance, C (F g⁻¹), was calculated according to

$$C = \frac{2I}{(dV/dt)m}$$

where I is the current (A), dV/dt is the slope of the discharge curve (V s⁻¹), and m is the mass (grams) of carbon in each electrode.

To perform high temperature electrochemical testing, the model capacitors were kept at 60 °C using an environmental test chamber (Tenney Environmental, Thermal Products Solutions, USA).

Supporting Information

Supporting Information is available from the Wiley Online Library or from the author.

ACKNOWLEDGMENTS

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