

Effect of structural differences of carbon nanotubes and graphene based iridium-NHC materials on the hydrogen transfer catalytic activity

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ABSTRACT

A proper design of heterogeneous molecular catalysts supported on carbon materials requires a systematic study of “metal–carbon support interactions” and their influence on catalytic activity. In this study, hybrid materials containing covalently anchored iridium N-heterocyclic carbene (NHC) organometallic complexes have been successfully prepared from oxidized and partially reduced carbon nanotubes (CNTs). The preparation method for these supported materials relies on the selective functionalization of the superficial hydroxylic groups using the imidazolium salt, 1-(3-hydroxypropyl)-3-methyl-1*H*-imidazol-3-ium chloride. The hydrogen transfer catalysis activity of these nanotube-based hybrid catalysts was tested by the reduction of cyclohexanone to cyclohexanol with 2-propanol, and the results of the tests were compared with those obtained using similar hybrid graphene-based catalysts. While EXAFS analysis revealed a common first coordination shell of the iridium atom for all the hybrid materials examined, independently of whether they were either supported on carbon nanotubes or graphene materials, catalytic activity in all the reduced materials was significantly superior. Moreover, catalytic systems based on reduced

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CNTs exhibited a better performance than those based on reduced graphene materials. Both these facts suggest there is a positive correlation between hydrogen transfer catalytic activity, reconstruction of the aromatic carbon structure and the smaller amount of oxygen functional groups.

1. Introduction

Technological development has provided us with available new one, two and three –dimensional carbon materials with a specific spatial and geometrical structure [1,2]. These highly versatile carbon materials are being currently used not only in electronic devices [3,4,5,6] but also as catalyst supports since their graphitic structure makes them inert in many chemical media [7]. For these reasons, in addition to their use as carbocatalysts, where oxidized functionalization is generally required [8,9,10], they have been profusely used as supports for hybrid materials having metal oxides, nanoparticles or molecular catalysts as the active sites [7,11,12,13,14]. These supports were initially evaluated as “innocent” materials due to their uniform hybridization, ordered structure and chemical inertness but it is known that the metal-support interactions are strongly dependent not only on the nature of the metal, but also on the type of carbon material chosen because their electronic [15,16], thermal [17] or mechanical properties [18,19] depend on their preparation method, origin and further modifications involving different surface treatments [20,21].

Carbon nanotubes (CNT) [22] and graphene [2] are the mono- and bi-dimensional long-range ordered versions, respectively, of these carbon materials. The synthetic methods used for carbon oxygen functionalization are well-established procedures. Graphene oxide (GO) is obtained from a severe oxidation process commonly employed for the chemical synthesis of graphene from graphite [23,24]. The sp^2 lattice of the graphene is disrupted upon oxidation by the introduction of oxygen functional groups, predominantly OH and epoxy in the interior of the basal plane while carboxylic acids are formed at the edges, and also where there are morphological defects (carbon atom vacancies) [25]. On the other hand, the oxidation processes, what introduce oxygen functionalization in the nanotubes (namely CNTO), are currently used as a purification step after synthesis [26]. Both oxidized carbon nanomaterials, GO and CNTO, have been utilized previously as catalysts themselves or as catalyst supports by making use of the surface groups in order to immobilize the active species [27,28,29,30]. Thus, not only metallic nanoparticles, but also enzymes or organometallic complexes have been anchored to these oxidized carbon nanomaterials

using esterification/amidation reactions with oxygen groups. The activation of carboxylic acids with SOCl_2 [31] or carbodiimides [32], and the reactivity of hydroxylic groups towards silanes [33] are routinely employed in order to functionalize CNTO and GO with the catalyst precursors. In general, the catalytic studies conducted on nanohybrid catalyst based on CNT or graphene systems have shown promising results in terms of activity, selectivity, stability and/or cyclability compared to the homogeneous or heterogeneous traditional systems [34].

Recently, our group found a protocol to immobilize iridium N-heterocyclic carbene (NHC) complexes on GO and thermally reduced GO materials by taking advantage of the carbon-based organic chemistry of the surface hydroxylic groups instead of using traditional siloxane chemistry [35]. The hybrid graphene–iridium–NHC materials were found to be efficient hydrogen-transfer catalysts exhibiting good recyclability. The effectiveness of the Ir(I)-NHC complexes as homogeneous catalysts in transfer hydrogenation of unsaturated compounds [36,37,38,39,40,41], in particular, those with hemilabile O- and N- donor functions in the NHC carbene ligand [42,43], is well known in the literature. In addition, an increase in the transfer hydrogenation catalytic activity of Ir-NHC complexes based on hemilabile O- and N- donor functions supported on CNT has been observed. [44] However, to date no systematic comparison of the catalytic activities and selectivities of different carbon-material based catalytic systems has been carried out and, to the best of our knowledge, the present study is unique in trying to establish the possible influence of the structure, surface chemistry and the reconstruction of the aromatic network on the catalytic activity of a hybrid organometallic complex–CNTO/GO carbon nanomaterial system. Despite the similar structure and properties of both materials, the geometrical and spatial differences could give rise to different local environments thereby influencing the catalytic activity.

This work has to tackle two critical points. Firstly, the covalent functionalization of oxidized and thermally reduced carbon nanotubes through surface -OH groups to produce iridium-supported NHC-CNTO complexes comparable to hybrid iridium NHC-GO catalysts has to be performed. Secondly, study the catalytic activity in the transfer hydrogenation of cyclohexanone with 2-propanol to compare with that of hybrid GO catalysts materials. Specific features of the structure and the surface chemistry of the hybrid catalysts have been used in this work to explain the differences in the catalytic performance.

2. Experimental

2.1. Materials

All the chemicals, including the graphite powder and multiwalled carbon nanotubes, were purchased from Aldrich. Reagent or HPLC grade was employed in all the experiments. Solvents were distilled immediately prior to use with appropriate drying agents or obtained from a Solvent Purification System (Innovative Technologies).

The oxidized CNTs (**CNTO**) utilized in this work were prepared by acid treatment of commercial bundle-type CNTs, as described previously [45]. The **GO** utilized in this work was prepared by applying a modified Hummers method to the commercial graphite, as reported in detail elsewhere [46]. **TRGO** and **TRCNTO** were obtained from **GO** and **CNTO** respectively by thermal treatment at 400°C in a horizontal furnace, under a nitrogen flow of 50 cm³ min⁻¹. The residence time at the final temperature was 60 min [46].

The imidazolium salt [MeImH(CH₂)₃OH]Cl (**1**) [47] and the compound [Ir(μ -OMe)(cod)]₂ [48] were prepared according to standard literature procedures.

2.2. Characterization of Nanotube and Graphene Materials and Hybrid Catalysts

Nuclear magnetic resonance (NMR) spectra were recorded on a Bruker Advance 400 spectrometer operating at 400.16 MHz (¹H). NMR chemical shifts are reported in ppm relative to tetramethylsilane and referenced to partially deuterated solvent resonances. Transmission electron microscopy (TEM) spectra were obtained on a JEOL 2000 EX-II instrument operating at 160 kV. Images of high-resolution transmission electron microscopy (HRTEM) were recorded using a JEOL JEM-2100F transmission electron microscope, equipped with a field-emission-gun (FEG) operating at 200 kV, and fitted with an Oxford Instruments microprobe to perform Energy-dispersive X-ray spectroscopy (EDX), in order to verify the atomic composition of the catalyst. The samples were prepared by casting a few drops of 1 mg mL⁻¹ ethanol suspensions of the materials over the carbon grids. To minimize exposure of the samples to the air, the suspensions were transferred to a lacey carbon grid using a glovebox filled with ultrahigh-purity argon and then to the TEM holder in order to minimize the time required to introduce them into the microscope. Elemental analyses were performed on a LECO-CHNS-932 micro-analyser equipped with a LECO-VTF-900 furnace coupled to a micro-analyzer. The X-ray photoemission spectroscopy (XPS) spectra were recorded

using a SPECS system operating under a pressure of 10^{-7} Pa and equipped with a Mg K α X-ray source. The functional groups in the graphene materials were quantified by deconvolution of the high resolution C1s XPS peak employing Gaussian and Lorentzian functions [49]. The binding energy profiles were deconvoluted as follows: undamaged structures of sp²-hybridized carbon (284.5 eV), damaged structures or sp³-hybridized carbons (and C-N groups, 285.5 eV), C-O (and C-N groups, 286.5 eV), C=O functional groups (287.7 eV) and COO groups at 288.7 eV. The amount of iridium present in the samples was determined by means of Inductively Coupled Plasma Mass Spectrometry (ICP-MS) in an Agilent 7700x instrument. The samples were digested following a method described elsewhere [50]. Briefly, 30 mg of sample was treated with 5 mL of a mixture of concentrated nitric and hydrochloric acid (3:1 ratio) at 180 °C for 3 h under microwave irradiation.

X-ray absorption spectroscopy (XAS) measurements were recorded at the BL22-Claess beamline from ALBA synchrotron (Spain) and at the BM23 beamline from the European Synchrotron Radiation Facility (ESRF, France). The measurements were carried out in transmission mode on pellets with an optimized thickness at room temperature. The beam was monochromatized by a fixed-exit-offset Si (311) double crystal in ALBA and a Si (111) one in ESRF. A harmonic rejection better than 10^{-5} was achieved by using a Si mirror coating on a dual toroid focusing mirror installed after the monochromator in both cases. The energy resolution $\Delta E/E$ was estimated to be about 8×10^{-5} at the Ir L₃-edge. A reference of metallic iridium was simultaneously measured for the calibration of the energy. XAS spectra of metal Ir, the complex [IrCl(cod)(MeImH(CH₂)₃OH)] and the salt IrCl₃ were also recorded as references for Ir(0), Ir(I) and Ir(III) respectively. The measurements obtained from both synchrotrons are comparable. Extended X-ray absorption fine structure (EXAFS) spectra were extracted using Athena software (part of the Demeter package) [51]. The Fourier transforms (FTs) of the EXAFS signals were calculated over a k-range of 3.0–13.3 Å⁻¹ using a sine window. The EXAFS structural analysis was performed using theoretical phases and amplitudes calculated by means of the FEFF-6 code [52] and fits to the experimental data were carried out in R-space with the help of ARTEMIS program (also part of Demeter package) [51].

2.3. Functionalization of Nanotube Materials with 1-(3-hydroxypropyl)-3-methyl-1H-imidazol-3-ium chloride (1)

Using a similar procedure to that recently reported for graphene materials [35], **CNTO** and **TRCNTO** were functionalized with the imidazolium salt, 1-(3-hydroxypropyl)-3-methyl-1H-

imidazol-3-ium chloride, [MeImH(CH₂)₃OH]Cl (**1**), following a two-step procedure. Briefly, **CNTO** or **TRCNTO** (100 mg) was dispersed in 20 mL of dichloromethane (DCM). The resulting dispersion was cooled to 0 °C in an ice bath and then, *p*-nitrophenylchloroformate (3.0 g, 15 mmol) and triethylamine (2.1 mL, 15 mmol) were added under an inert atmosphere. The mixture was stirred for 24 h and allowed to reach room temperature slowly. The resulting solids were filtered and washed three times with DCM (20 mL) and then dried under vacuum for 2 h. In a second step, the imidazolium salt **1** (100 mg, 0.560 mmol) and a catalytic amount of triethylamine (0.2 mL) were added, under an inert atmosphere, to the dispersions of nanotube solids obtained in step 1 in tetrahydrofuran (THF) (15 mL), and refluxed for 24 h. The products were obtained by centrifugation/filtration, washed with THF (3 x 20 mL), DCM (3 x 20 mL), and ethanol (3 x 20 mL) and then, vacuum dried at 100 °C in a preheated furnace until constant weight. The nanotube samples obtained were labeled as **CNTO-1** and **TRCNTO-1** depending on the parent material used in each case, **CNT** or **TRCNTO**, respectively.

2.4. Preparation of Hybrid Catalysts CNTO-1-Ir and TRCNTO-1-Ir

CNTO-1 or **TRCNTO-1** (100 mg) were made to react with [Ir(μ -OMe)(cod)]₂ (100 mg, 0.150 mmol) (cod = 1,5-cyclooctadiene) in THF (10 mL) under an argon atmosphere. The mixtures were refluxed for 2 days and then immersed in an ultrasonic bath for 30 min at room temperature. The resultant solids were recovered by centrifugation, washed with THF (5 x 10 mL) and diethyl ether (2 x 5 mL), and dried under vacuum to produce **CNTO-1-Ir** and **TRCNTO-1-Ir**.

2.5. General Procedure for Transfer Hydrogenation Catalysis.

The catalytic transfer hydrogenation reactions were carried out under an argon atmosphere in thick glass reaction tubes fitted with a greaseless high-vacuum stopcock. In a typical experiment, the reactor was charged with a solution of cyclohexanone (0.52 mL, 5.0 mmol) in 2-propanol (4.5 mL), internal standard (mesitylene, 70 μ L, 0.5 mmol), base (0.1 mL, 0.025 mmol of a KOH solution 0.24 M in 2-propanol) and catalyst (0.005 mmol, 0.1 mol%). The weight of the supported catalysts used in each experiment was calculated according to ICP measurements, assuming that all the iridium in the sample corresponded to active catalyst sites, 9.34 mg of **GO-1-Ir** (10.2 %wt. of iridium) and 20.44 mg of **TRGO-1-Ir** (4.7 %wt of iridium) were used for the GOs, and 7.7 mg of **CNTO-1-Ir** (12.5 %wt. of iridium) and 9.5 mg of **TRCNTO-1-Ir** (10.1 %wt. of iridium) were

employed for the CNTs. The resulting mixture was stirred at room temperature until complete dissolution of the homogeneous catalyst, $[\text{IrCl}(\text{cod})(\text{MeIm}(\text{CH}_2)_3\text{OCOCH}_3)]$ (**Ir-ImidO**) [35], or for 10 min in the case of the heterogeneous catalyst, and then placed in a thermostatic oil bath at the required temperature, typically 80 °C. Conversions were determined by gas chromatography analysis (HP-INNOWax capillary column -0.4 μm , 25 m x 0.2 mm i.d.-) under the following conditions: a column temperature of 35 °C (2 min) up to 220 °C at 10 °C min^{-1} and a flow rate of 1 $\text{cm}^3 \text{min}^{-1}$ using ultrapure He as carrier gas.

Once the reaction was completed, the hybrid catalysts were recovered by centrifugation and washed with additional amounts of 2-propanol (3x10 mL). Several catalytic cycles were performed with these materials, under the same experimental conditions, without adding any fresh catalyst precursor. The last cycle was carried out without an inert atmosphere.

3. Results and discussion

3.1. Preparation and catalytic activity of Nanotube-based Hybrid Catalysts

The nanotube-based hybrid catalysts were prepared following the reaction path depicted in Fig. 1 [35]. The “isolated” hydroxyl groups in the nanotubes react with *p*-nitrophenylchloroformate, with the subsequent formation of the corresponding *p*-nitrophenyl carbonate esters [53,54]. Thereafter, the nucleophilic OH-ending group of the imidazolium salt $[\text{MeImH}(\text{CH}_2)_3\text{OH}]\text{Cl}$ (**1**) easily displaces the *p*-nitrophenol group, resulting in the formation of the carbonate intermediates, **CNTO-1** and **TRCNTO-1**. Reaction of the imidazolium groups of these materials with the methoxo iridium(I) dimer compound $[\text{Ir}(\mu\text{-OMe})(\text{cod})]_2$ (cod = 1,5-cyclooctadiene) gave the nanotube-based hybrid materials **CNTO-1-Ir** and **TRCNTO-1-Ir**.

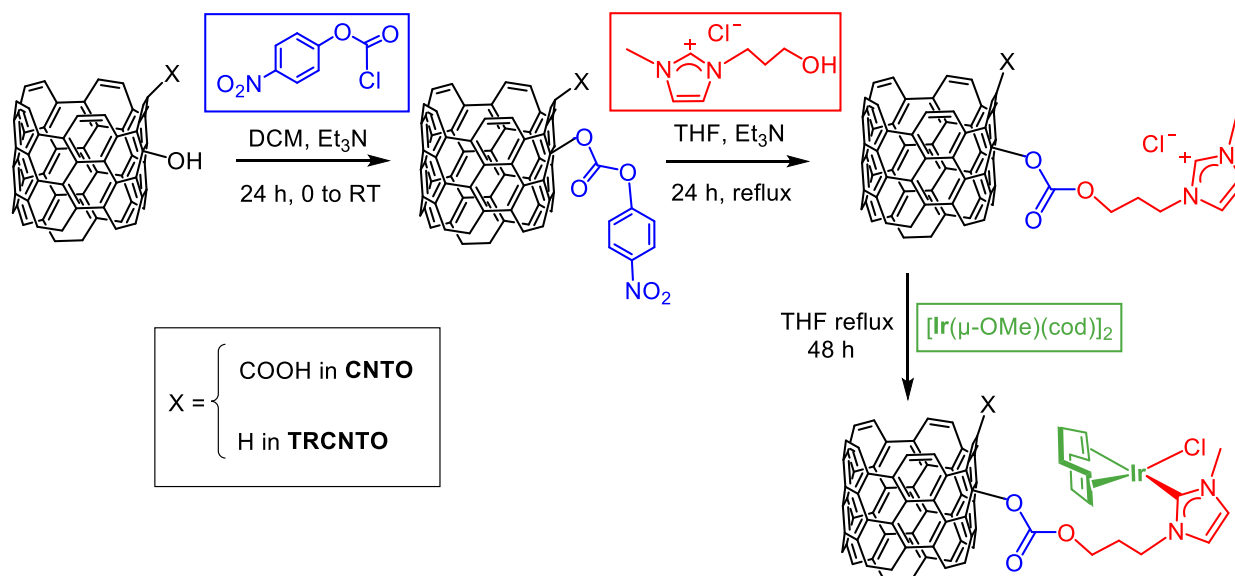


Fig. 1 - Synthesis of hybrid nanotube-based-Ir-NHC materials **CNTO-1-Ir** and **TRCNTO-1-Ir** through the covalent functionalization of the parent nanotube materials, **CNTO** and **TRCNTO**, with the imidazolium salt [MeImH(CH₂)₃OH]Cl (**1**).

The covalent linkage of the imidazolium ligand to the nanotube was confirmed by the characterization of the intermediate species **CNTO-1** and **TRCNTO-1**. These materials still form stable suspensions in acetone (see Supporting Information) and their ¹H NMR (acetone-d₆) spectra show a set of signals typical of imidazolium groups, i.e., 7.41/7.35 ppm (H4 and H5) and 8.64 ppm (H2) (see Supporting Information). Neither the signals corresponding to the nitrophenyl fragment nor those corresponding to the imidazolic -OH group can be observed in the spectra. The presence of nitrogen and chlorine in their XPS spectra at a N:Cl ratio of 2 also confirms these results (Table 1). From these analyses it is estimated that almost four times more imidazolium ligand is introduced in **CNTO-1** than in **TRCNTO-1**, what could be justified considering a larger amount of OH groups (at basal planes) in the **CNTO** parent material compared to those of **TRCNTO** as confirmed by FTIR analysis, together a better dispersion capacity of the material in the polar reaction medium (see Supporting Information). Additionally, the larger percentage of the C-O band on the C1s XPS curve, Table 1, observed for **TRCNTO** is attributed to an increase in unreactive epoxy or ether groups [55]. Similar tendencies were observed in the functionalization of graphene oxides and thermally reduced graphene oxides under the same experimental conditions [35].

Table 1 - Elemental analysis and XPS characterization of imidazolium-functionalized carbon materials.

Sample	Elemental Analysis		XPS (atom %)						
	<i>C/O</i>	<i>N (wt.%)</i>	<i>N</i>	<i>Cl</i>	<i>C_{sp}²</i>	<i>C_{sp}³</i>	<i>C-O</i>	<i>C=O</i>	<i>COO</i>
CNTO	3.8	0.0	0.0	0.0	64.4	17.1	7.0	3.2	8.2
TRCNTO	11.8	0.0	0.0	0.0	70.6	13.9	9.3	2.3	3.9
GO	1.0	0.1	0.0	0.0	36.2	8.4	29.6	17.8	7.9
TRGO	2.9	0.0	0.0	0.0	71.4	9.9	11.5	4.8	2.4
CNTO-1	1.7	5.6	4.7	2.0	53.7	20.9*	18.8*	4.7	1.8
TRCNTO-1	6.7	2.5	1.4	0.6	57.1	17.8*	16.4*	7.2	2.0
GO-1	1.4	6.9	5.6	2.5	42.7	24.0*	21.1*	6.2	6.0
TRGO-1	3.8	1.4	1.4	0.6	58.2	15.6*	14.4*	6.1	6.9

* Includes the C-N bonds

The HRTEM images in Figure 2 of the hybrid materials with the N-heterocyclic carbene (NHC) iridium complexes show homogeneous distribution of electron-dense regions associated with the supported iridium species. In spite of the presence of these anchored iridium species in **CNTO** and **TRCNTO**, they exhibit the same appearance in terms of number of layers and interlayer distances as their parents (see Supporting Information) which indicates that functionalization has not caused any damage to the nanotube layers. The iridium spots observed (also confirmed by EDX, see Supporting Information) exhibit diameters from 0.17-0.27 nm (highlighted with white circles), which are in the range of the molecular complexes, to 1.2-1.4 nm and even larger, which are the case of clusters or nanoparticles possibly formed during beam irradiation inside the microscope chamber. Similar size distributions were observed for the graphene-based hybrid catalyst [35] or even for other supported molecular iridium catalysts [56,44].

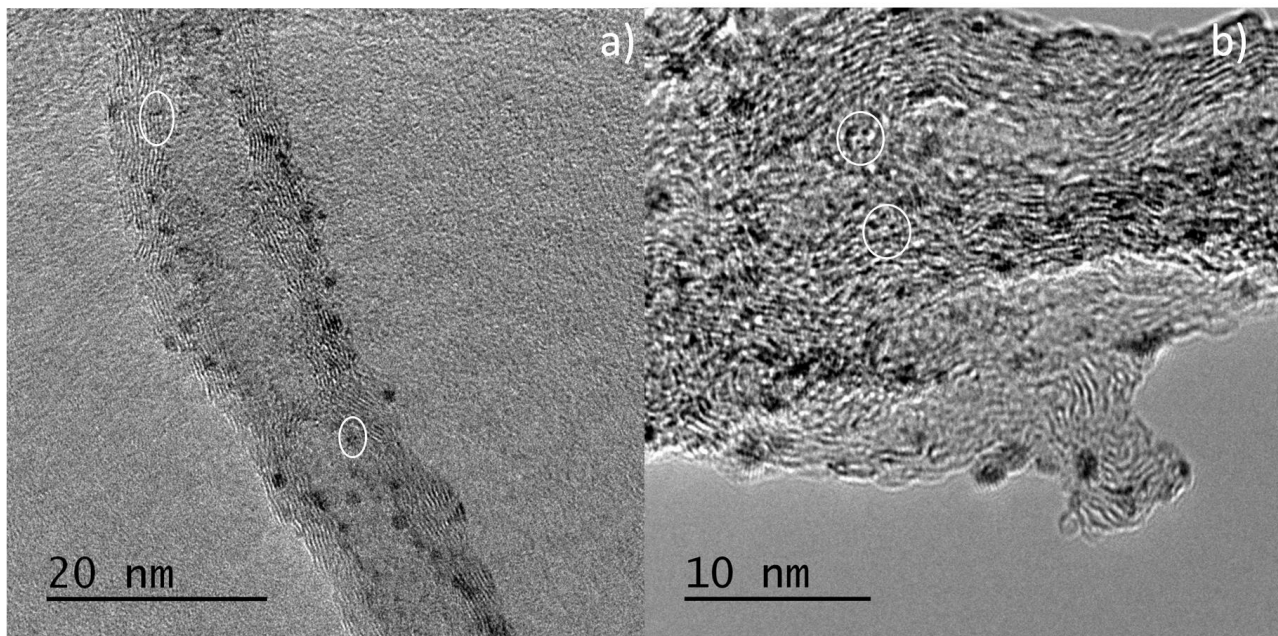


Fig. 2 - HRTEM images of a) **CNTO-1-Ir** and b) **TRCNTO-1-Ir**.

The amount of iridium in the samples, calculated by means of ICP-MS, was 12.5 and 10.1 wt.% for **CNTO-1-Ir** and **TRCNTO-1-Ir**, respectively. The percentage of iridium in the **TRCNTO-1-Ir** material lies within the expected range, as is confirmed by the amount of imidazolium salt linked to the material (98%). However, the metal content in **CNTO-1-Ir** is below the estimated value (52%), which might be explained by the reduced amount of imidazolium ligands available due to steric hindrance resulting from the high density of functional groups [32].

The hydrogen-transfer catalytic activity of the samples was studied in the reduction of cyclohexanone to cyclohexanol using 2-propanol both as hydrogen source and as a non-toxic solvent. The catalytic conditions used were catalyst loads of 0.1 mol %, with 0.5 mol % of KOH as co-catalyst and 80 °C, i.e., identical conditions to those recently reported for the same catalyst supported on graphene materials [35]. For comparative purposes, these results are also depicted in Fig. 3 and summarized in Table 2, together with those of the molecular acetoxy-functionalized NHC complex $[\text{IrCl}(\text{cod})(\text{MeIm}(\text{CH}_2)_3\text{OCOCH}_3)]$ (**Ir-Imido**) [44]. It is worth noting that none of the iridium free nanotube and graphene materials showed any catalytic activity (see Supporting Information). Additionally, the nanotube and graphene based catalysts prepared by means of the same procedure but without NHC linkages showed no catalytic activity at all after the first cycle probably due to leaching of the iridium after washing (see Supporting Information).

Table 2 - Catalytic Hydrogen Transfer from 2-Propanol to Cyclohexanone using carbon nanotube and graphene based iridium-NHC hybrid catalysts, and the related acetoxy-functionalized NHC-Ir homogeneous catalyst [IrCl(cod)(MeIm(CH₂)₃OCOCH₃)] (**Ir-ImidO**).^{a,b}

Catalyst	Time (min)	Conversion (%)	TOF ₀ (h ⁻¹) ^c	TOF ₅₀ (h ⁻¹) ^d	Recyclability	
					Mean Conversion 4 cycles-Argon (%)	Conversion 5 th cycle on Air (%)
CNTO-1-Ir	210	89	3000	1304	93	92
TRCNTO-1-Ir	100	92	6000	3000	93	93
GO-1-Ir	840	93	1020	375	92	94
TRGO-1-Ir	150	90	3600	1685	93	92
Ir-ImidO	220	92	3000	1485	-	-

^a Reaction conditions: catalyst/substrate/KOH ratio of 1/1000/5, 0.1 mol% of catalyst in 2-propanol (5 mL) at 80 °C. ^b The reactions were monitored by GC using mesitylene as internal standard. ^{c, d} TOF, turnover frequency [(mol product/mol catalyst)/time (h)], were calculated at initial time (60s) or at 50% of conversion, TOF₀, and TOF₅₀, respectively.

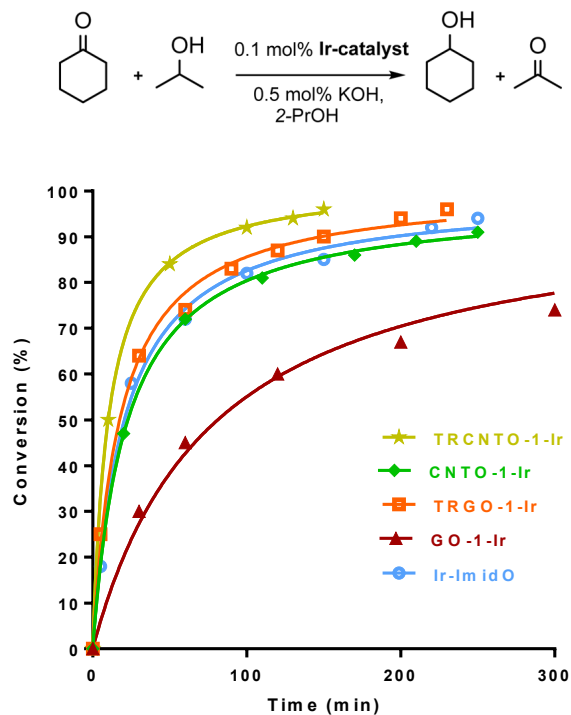


Fig. 3 - Reaction profiles showing the transfer hydrogenation of cyclohexanone by carbon nanotube and graphene based iridium-NHC hybrid catalysts and the related homogeneous molecular catalyst $[\text{IrCl}(\text{cod})(\text{MeIm}(\text{CH}_2)_3\text{OCOCH}_3)]$ (**Ir-ImidO**).

Taking into account the carbon nanotube-iridium-NHC catalysts, the catalytic activity of the partially reduced materials **TRCNTO-1-Ir** is higher than that of **CNTO-1-Ir**, leading to yields higher than 90% in 100 min instead of 210 min required for the fully oxidized catalyst (conversions determined by GC with mesitylene as internal standard). A similar trend was observed for the graphene-based hybrid catalysts, which confirms that for this type of catalytic reactions, a partial reduction of the nano-support has a positive effect on the catalytic activity. Additionally, all iridium-NHC carbon nano-supports can easily be recycled by a method consisting of simple centrifugation, washing of the catalysts with fresh 2-propanol (4 x 5 mL) and addition of more cyclohexanone, KOH and 2-propanol. The recycling processes render conversions above 90 % after five catalytic runs with reaction times of 2.5 h for **TRGO-1-Ir** and 13 h for **GO-1-Ir** for each run, while the catalytic reactions take 1.5 h and 4 h for **TRCNTO-1-Ir** and **CNTO-1-Ir**, respectively. All the reaction profiles are similar in the successive experiments, even in the case of the last cycle, where there is no protective atmosphere (see Supporting Information). In addition, ICP-MS

measurements of the reused materials after the fifth catalytic runs showed similar iridium content (12.3 and 10.0 wt.% for **CNTO-1-Ir** and **TRCNTO-1-Ir**, respectively) than that of the fresh catalysts (12.5 and 10.1 wt.%, respectively) which reveal no leaching of the iridium on those materials.

Of special interest, however, is the fact that the catalytic activity of all the carbon nanotube-based hybrid catalysts is greater than that of their corresponding graphene material. In this regard it is worth noting that the catalytic activity of **GO-1-Ir** is also lower than that of the homogeneous catalyst. In order to explain these findings, a detailed characterization of the supported catalyst was performed.

3.2. Structural features of the nanotubes and graphene-NHC-Iridium hybrid catalysts

In order to gain an insight into the local structure of the iridium atoms in the hybrid catalysts, XAS measurements were performed at room temperature. All the XANES spectra show similar features (see Supporting Information) and their edge position is equal to that observed in the Ir(I) reference compound, $[\text{IrCl}(\text{cod})(\text{MeIm}(\text{CH}_2)_3\text{OH})]$, and therefore intermediate between the metal Ir and IrCl_3 indicating a similar state of oxidation of (I) for the Ir atoms in these hybrid catalysts (see Supporting Information). This is in agreement with the high-resolution Ir4f XPS band obtained for all hybrid nanotube and graphene based iridium catalysts (Fig. 4) which shows two peaks (corresponding to $\text{Ir}4f_{7/2}$ and $\text{Ir}4f_{5/2}$) centered at 62.4 and 65.6 eV, that are characteristic of iridium(I) compounds [57].

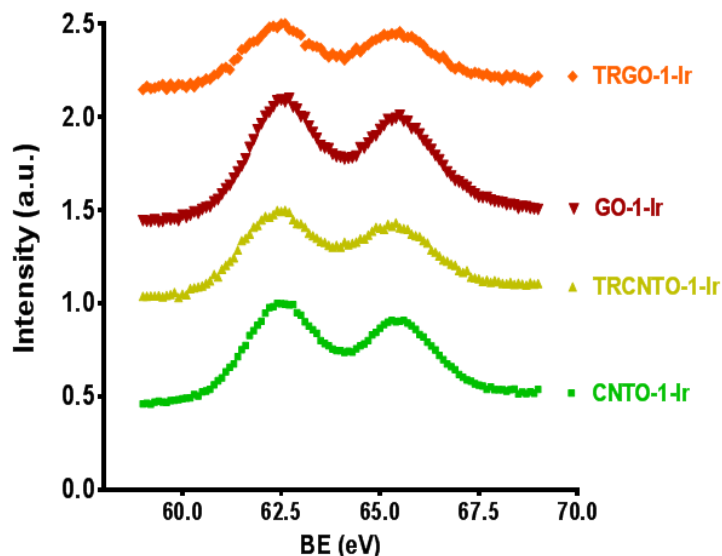


Fig. 4 - High-resolution Ir4f XPS spectra obtained for the all the hybrid carbon nanotube and graphene based iridium-NHC catalysts.

Fig. 5a presents the k^2 -weighted EXAFS signals for all studied hybrid catalysts and the reference compound [IrCl(cod)(MeIm(CH₂)₃OH)] for the sake of comparison. The four hybrid catalysts show similar oscillations in the spectra though a higher damping is observed in the signal of **GO-1-Ir**. However, the signal for the reference compound shows a clear interference at $k \sim 8-11 \text{ \AA}^{-1}$. In order to highlight the differences between the compounds under study, Fig. 5b displays the modulus of the FTs of the previous EXAFS spectra to provide a measure of the pseudo radial distribution function around the Ir atom. The hybrid catalysts show a first coordination shell composed by a single peak at $R \sim 1.7 \text{ \AA}$ (without phase shift correction) while a second peak or shoulder is observed in the reference complex close to 2.1 \AA (related to the above mentioned interference) that it is ascribed to the Ir-Cl bond. Then, the first main isolated peak in the FTs of the catalysts corresponds to the first-neighbors coordination shell, which is only composed of light elements, likely ascribed to C (and/or some O). The intensity of this peak is significantly smaller for the **GO-1-Ir** samples than in the other three compounds. The decrease in intensity is related to the high damping observed in its EXAFS signal and indicates greater structural disorder. Structures beyond 2.2 \AA correspond to mixed contributions from the next-neighbors coordination shells. Although all the samples exhibit different patterns in this range, the signals are significantly weaker and broader, which prevent them from being analyzed. Therefore, EXAFS analysis was limited to the first coordination shell, performing the study between 1.15 and 2.25 \AA in the R-space.

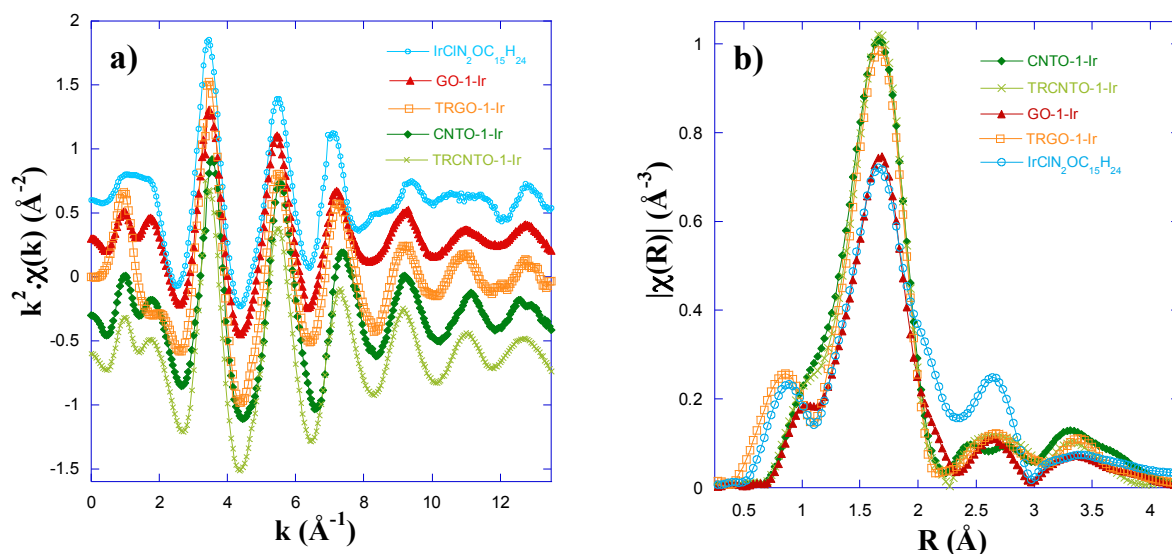


Fig. 5 - a) k^2 -weighted EXAFS spectra for the four hybrid catalysts and the [IrCl(cod)(MeIm(CH₂)₃OH)] reference compound (IrClN₂OC₁₅H₂₄). Data are shifted in the vertical scale for the sake of comparison. b) Moduli of the FTs of the $k^2\chi(k)$ EXAFS signals for the same samples.

In order to obtain an accurate model for the analysis of our EXAFS data, we have drawn on our previous experience [43]. The first coordination shell of the Ir atom in the reference complex, [IrCl(cod)(MeIm(CH₂)₃OH)], is composed of a Cl atom, C number 1 (C₁) in the imidazol-2-ylidene ring and 4 C atoms of both C=C bonds present in the cyclooctadiene ligand (see Fig. 6a).

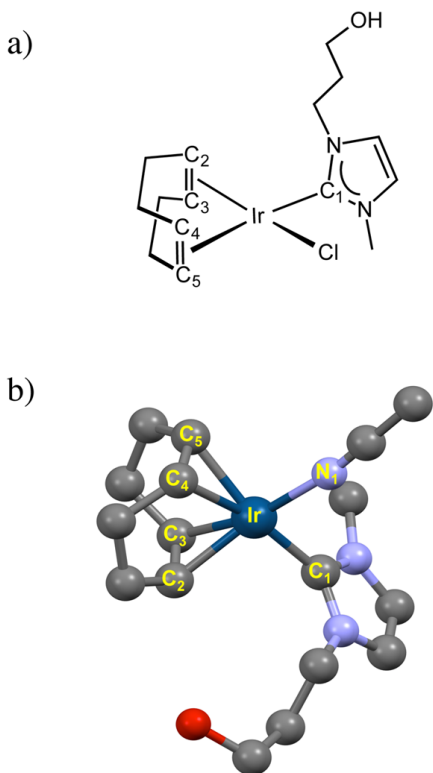


Fig. 6- a) Schematic molecular structure of the reference neutral chloride compound $[\text{IrCl}(\text{cod})(\text{MeIm}(\text{CH}_2)_3\text{OH})]$. b) Molecular structure of the cation of $[\text{Ir}(\text{NCCH}_3)(\text{cod})(\text{MeIm}(\text{CH}_2)_3\text{OH})][\text{BF}_4]$ (Hydrogen atoms have been omitted for clarity). Selected bond lengths (\AA) and angles ($^\circ$): Ir–C(1) 2.031(3) , Ir–C(2) 2.127(3), Ir–C(3) 2.134(4), Ir–C(4) 2.199(3), Ir–C(5) 2.186(3), Ir–N(1) 2.032(3), C(1)–Ir–N(1) 92.44(12), C(1)–Ir–C(2) 90.58(13), N(1)–Ir–C(5) 87.55(12), C(2)–Ir–C(4) 90.11(14).

The hybrid materials should have, in principle, the IrCl(cod) metal fragment anchored to the carbon matrix (either **CNTO** or **GO**) through the NHC linker (Fig. 1). However, having in mind a possible iridium-support interaction, the ionization of the chlorido ligand cannot be discarded. Thus, replacement of the Cl by another anion, probably an O atom from the oxidized carbon matrix, could be possible. As shown in Fig. 5, the EXAFS signal rules out the presence of Cl atoms in the first coordination shell of these hybrid catalysts (lack of interference at $k \sim 8\text{--}11\text{\AA}^{-1}$), as reported in related compounds [35,44]. Therefore, our proposed model for the first coordination shell surrounding Ir atoms inside the catalysts is composed of five Ir-C and one Ir-O contributions. The Ir-C and Ir-O phase shifts and back scattering amplitudes were calculated from the square planar

reference compound $[\text{Ir}(\text{NCCH}_3)(\text{cod})(\text{MeIm}(\text{CH}_2)_3\text{OH})][\text{BF}_4]$ and their bond lengths were taken as the starting point (see Fig. 6b). The distances of Ir to the carbons in the diolefin range between $R_2 = 2.13 \text{ \AA}$ and $R_3 = 2.19 \text{ \AA}$ and they are coupled in pairs (2+2). The bond length to the C_1 is $R_1 \sim 2.03 \text{ \AA}$ and the Ir-O bond length is assumed to be close to this value. As a result, the iridium centre is bonded to six light elements with the following distribution: $R_1 (\times 2)$, $R_2 (\times 2)$ and $R_3 (\times 2)$. A total of 4 parameters are refined, with an average inner potential correction of the threshold (ΔE_0), an average Debye-Waller factor (σ^2) and two distances, R_2 and R_3 are coupled with the same free parameter. The amplitude reduction factor S_0^2 is fixed to 1 in agreement with the value obtained for the reference compounds. Fig. 7 compares the best fit and experimental spectra corresponding to the moduli of the FTs of the k^2 -weighted EXAFS (Fig. 7a) with the corresponding Fourier filtered spectra in k -space (Fig. 7b). There is considerable agreement in spite of the simplicity of our structural model.

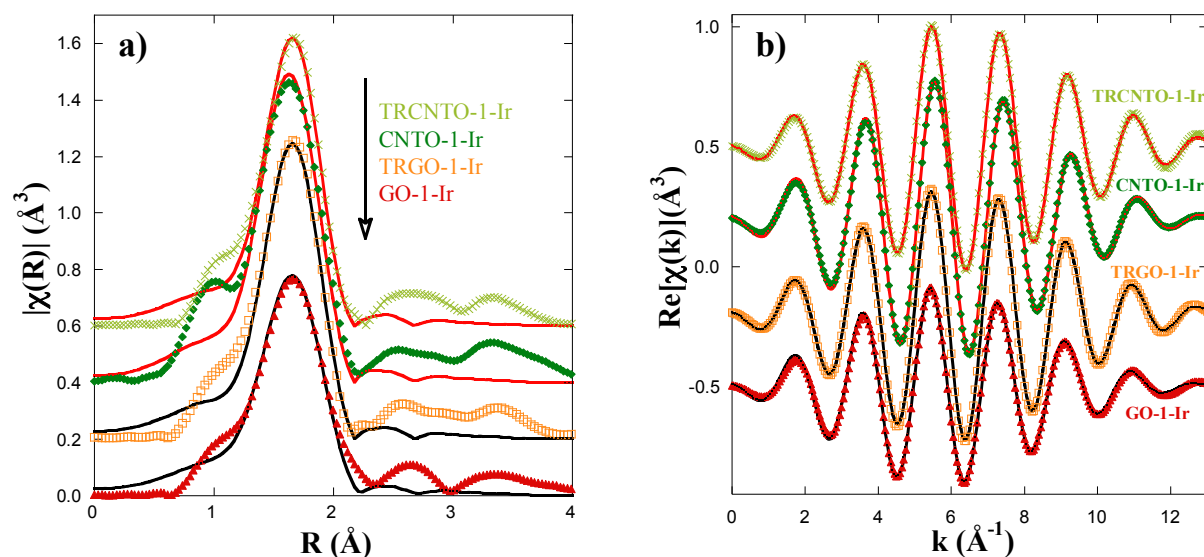


Fig. 7 - Comparison between experimental data (circles) and best fits (solid lines) for the spectra of the hybrid catalyst. a) Moduli of the FTs of the $k^2\chi(k)$ EXAFS signals, and b) real part of the Fourier-filtered spectra in k -space corresponding to the first shell coordination.

The relevant structural parameters obtained for the four hybrid catalysts are summarized in Table 3. The distribution of bond lengths is very similar for the four materials and they are slightly shorter than that of the reference compound. The only significant difference concerns the **GO-1-Ir**

compound that exhibits greater local disorder and this is reflected in a larger σ^2 -value. This disorder is also reflected in the higher standard deviation of the refined distances and may indicate a broader distribution of Ir-C paths.

Table 3 - Best fit structural parameters for the first coordination shell of the hybrid catalysts at the Ir L₃-edge: Average inner potential correction (eV), bond lengths (Å), average Debye-Waller factor (Å²) and a residual factor of the fit.^a

Sample	ΔE_0 (eV)	R_1 (Å)	R_2 (Å)	R_3 (Å)	$10 \cdot \sigma^2$ (Å ²)	R_F
CNTO-1-Ir	8.3(7)	2.03(2)	2.12(2)	2.18(2)	3.3(4)	0.0029
TRCNTO-1-Ir	7.7(7)	2.02(2)	2.10(2)	2.16(2)	3.0(3)	0.0028
TRGO-1-Ir	8.2(9)	2.02(2)	2.11(2)	2.17(2)	3.6(9)	0.0049
GO-1-Ir	8.5(8)	2.04(4)	2.12(4)	2.18(4)	7.0(22)	0.0085

^a The residual factor accounts for the misfit between the actual data and the theoretical calculations [51]. Numbers in parentheses are the errors estimated from different analyses to the best significant digit.

If these results are compared with the catalytic activity exhibited by these materials it will be observed that **GO-1-Ir** shows by far the lowest catalytic activity (Fig. 3, Table 2). This reflects the stronger local disorder of the Ir first-neighbors coordination shell in this sample that could be due to the high oxidation degree of this support (Table 1) has a negative effect on the catalytic activity of the compound. The catalytic results are however also determined by the peculiarities of the particular carbon support. It can be inferred from this results that the next-neighbors coordination shells, which are different in all the samples studied, also plays a crucial role in the catalytic system. Unfortunately, the strong disorder inferred from the EXAFS spectra prevents an accurate analysis of these coordination shells. Therefore, in order to gain more information about the “local environment” surrounding the active Ir centers in the hybrid catalyst, the XPS C1s analyses of the materials have been closely examined. The results are summarized in Fig. 8.

Similar profiles were obtained for samples **TRCNTO-1-Ir** and **TRGO-1-Ir**, which are also similar to that of **CNTO-1-Ir**. As the first-neighbors coordination shell of the Ir supported on these materials is similar, the differences in the catalytic activity must be related to the differences in the

surroundings resulting from the carbon support. Interestingly, the catalytic activity of these three samples, following the order **TRCNTO-1-Ir** > **TRGO-1-Ir** > **CNTO-1-Ir** (Fig. 3), is enhanced with the increase in the amount of Csp² bonds and the decrease in functional groups in the supports, which suggests that, for the catalysts with similar first-neighbors coordination shells, a more aromatic support with fewer defects would be favorable for the catalyst performance in the process under study. On the other hand, **GO-1-Ir** also shows the highest percentage of sp³ carbon bonds and oxygen functional groups (including the C-O, C=O and COO groups). The poorer catalytic activity of this sample (Fig. 3) is therefore probably related to the strong local disorder in the graphene layer supporting the Ir catalyst. However, the extent of the disorder and its proximity to the Ir (particularly of the C-O and C=O derived functional groups in the basal planes) also seem to affect the first-neighbors coordination shell, (as can be observed from the EXAFS curves, Fig. 5 and 7) resulting in a drastic decrease in the catalytic activity.

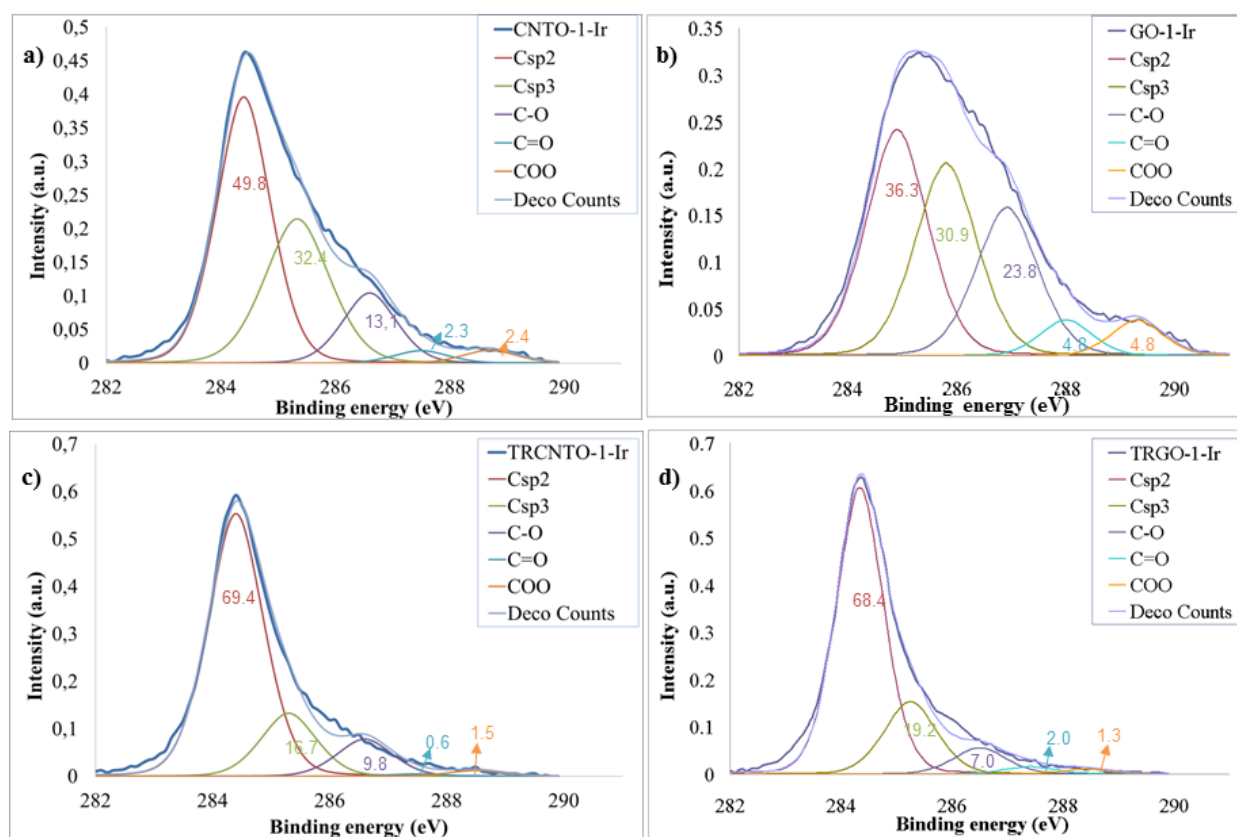


Fig. 8 - Deconvolution results of the XPS C1s curves of a) **CNTO-1-Ir**, b) **GO-1-Ir**, c) **TRCNTO-1-Ir** and d) **TRGO-1-Ir**.

4. Conclusions

The protocol for functionalizing oxidized carbon nanomaterials through their surface OH groups on the basis of the activation with *p*-nitrophenylchloroformate can be successfully applied to oxidized and partially reduced carbon nanotubes in order to immobilize iridium NHC-complexes in their outer and inner walls. The carbon nanotube based iridium-NHC hybrid materials were efficient catalysts for the reduction of cyclohexanone by transfer hydrogenation. These hybrid catalysts have shown a superior catalytic performance than related graphene-based hybrid materials, prepared by the same functionalization strategy, in both the oxidized and thermally reduced graphene supports.

The study of the local structure of the iridium atoms in the hybrid catalyst by EXAFS has revealed the replacement of the chlorido ligand by an O atom from the oxidized carbon matrix as a result of the iridium-support interaction, which might be facilitated by the presence of a flexible linker to carbon matrix. On the other hand, as it has been shown by XPS studies, the greater number of structural defects in the graphenic based materials than in their respective nanotubes might explain their different performances. Although the same first-neighbors coordination shell is formed in every hybrid catalyst, the strong structural disorder in this very local environment leads to very poor results, as in the case of the **GO-1-Ir** catalyst. In contrast, the high degree of order in the sp^2 lattice, combined with an appropriate amount of oxygen groups, provides a stability that enhances the catalytic activity, as in the case of **TRCNTO-1-Ir**.

Supporting Information.

Supplementary data associated with this article can be found, in the online version, at <http://dx.doi.org/10.1016/j.carbon.xxxxxxx>.

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